

Name: _____ Date: _____ Period: _____

Periodic Trends Practice Worksheet

1. What is the difference between electronegativity and ionization energy?
2. Why does fluorine have a higher ionization energy than iodine?
3. Why do elements in the same family generally have similar properties?
4. Which is the largest atom in Group 4a (14)?
5. Which is the smallest atom in Group 7a (17)?
6. Which is the smallest atom in period 5?
7. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
8. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
9. Arrange the following atoms in order of decreasing atomic radius. Na Al P Cl Mg
10. For each of the following pairs, circle the element that is larger?
a) N^{3-} or F^{-1} b) Mg^{2+} or Ca^{2+} c) Fe^{2+} or Fe^{3+}
11. Circle the element in each pair has the larger radius?
a) Mg or Mg^{2+} b) O or O^{2-} c) K^{+} or Cl d) P^{3-} or S^{2-}
12. In each of the following pairs, circle the species with the higher first ionization energy:
a) Li or Cs (b) Cl or Ar (c) Ca or Br (d) Na or Ne (e) B or Be
13. In each of the following pairs, circle the species with the larger atomic radius:
a) Mg or Ba (b) S or S^{2-} (c) Cu^{+2} or Cu (d) He or H (e) Na or Cl
14. Circle the best choice in each list:
a) highest first ionization energy: C, N, Si
b) highest electronegativity: As, Sn, S
c) largest radius: S^{2-} , Cl^{-} , Cl
15. Order the following groups from largest to smallest radii.
a) Ar, Cl, K, S^{2-}
b) C, Al, F, Si
c) Na, Mg, Ar, P
d) I^{-} , Ba^{2+} , Cs^{+1} , F^{-1}

Chemistry Practice Test Periodic Trends And Orbitals

Claire Aldridge, Karl Lee



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Chemistry Class 11 Dr. S C Rastogi, Er. Meera Goyal, 2022-06-14 This Book has been written in according with the New Syllabus of Madhyamik Shiksha Mandal Madhya Pradesh Bhopal based on the curriculum of CBSE NCERT Including solved questions of NCERT book based on new examination pattern and mark distribution Highly Useful for NEET AIIMS IIT JEE APJ AKTU and Engineering Medical Examinations Syllabus Unit I Some

Basic Concepts of Chemistry Unit II Structure of Atom Unit III Classification of Elements and Periodicity in Properties Unit IV Chemical Bonding and Molecular Structure Unit V States of Matter Gases and Liquids Unit VI Chemical Thermodynamics Unit VII Equilibrium Unit VIII Redox Reactions Unit IX Hydrogen Unit X s Block Elements Alkali and Alkaline earth metals Group 1 and Group 2 Elements Unit XI Some p Block Elements General Introduction to p Block Elements Unit XII Organic Chemistry Some Basic Principles and Techniques Unit XIII Hydrocarbons Classification of Hydrocarbons Unit XI V Environmental Chemistry Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table *Chemistry Class XI - SBPD Publications* Dr. Subhash Chandra Rastogi, ,Er. Meera Goyal,2021-10-16 Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table Cracking the GRE Chemistry Subject Test Princeton Review (Firm),2005 The GRE subject tests are among the most difficult standardized exams Rather than testing general problem solving skills they require highly specialized knowledge The experts at The Princeton Review have thoroughly research each subject test to provide students with the most thorough up to date information available Students don t need to relearn the entire histories of their fields just what they need to know to earn high scores on the exams Each guide includes one full length practice exam complete with comprehensive explanations for every solution *Kaplan SAT: Chemistry* Claire Aldridge,Karl Lee,2006-02 Includes 3 Full length practice tests Comprehensive review of all the tested material with practice quizzes for each chapter Effective strategies to maximize your score **Chemistry 2007** David Wilson,2006-12 Kaplan s guide includes 2 full length practice tests Diagnostic test to target areas for score improvement Detailed answer explanations Hundreds of practice questions from calculations of chemical equations to organic chemistry Explanations of important terms formulas and concepts Powerful strategies to help you score higher **Prentice Hall Chemistry** ,2000 *Fundamentals of Engineering* Merle C. Potter,1993 *General Chemistry* Kenneth W. Whitten,Raymond E. Davis,1996 Are you looking for the key to success in your chemistry class In CHEMISTRY you will find a strong molecular reasoning focus problem solving exercises and an innovative online homework management system that will prepare you for any challenge you might encounter The textbook is filled with learning aids that will help you master concepts of the course General Chemistry Ralph H. Petrucci,William S. Harwood,1993 *General Chemistry Principles and Modern Applications* is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous

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