

## ASSESSMENT

**Chapter Test B****Chapter: Chemical Bonding**

**PART I** On the line at the left of each statement, write the letter of the choice that best completes the statement or best answers the question.

- \_\_\_\_\_ 1. An ionic bond results from electrical attraction between  
a. cations and anions.  
b. atoms.  
c. dipoles.  
d. orbitals.
- \_\_\_\_\_ 2. A nonpolar covalent bond is unlikely when two atoms of different elements join because the atoms are likely to differ in  
a. density.  
b. state of matter.  
c. electronegativity.  
d. polarity.
- \_\_\_\_\_ 3. Bond length is the distance between two bonded atoms at  
a. their minimum potential energy.  
b. their maximum kinetic energy.  
c. their maximum potential energy.  
d. one-half the diameter of the electron cloud.
- \_\_\_\_\_ 4. To draw a Lewis structure, it is not necessary to know  
a. which atoms are in the molecule.  
b. bond energies.  
c. the number of valence electrons for each atom.  
d. the number of atoms in the molecule.
- \_\_\_\_\_ 5. For multiple covalent bonds to form in molecules, the molecules must contain carbon, nitrogen, or  
a. chlorine.  
b. hydrogen.  
c. oxygen.  
d. helium.
- \_\_\_\_\_ 6. The principle that states that atoms tend to form compounds in which each atom has eight electrons in its highest occupied energy level is called the  
a. rule of eight.  
b. configuration rule.  
c. Avogadro principle.  
d. octet rule.

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