

# REACTIVITY TRENDS IN THE PERIODIC TABLE

**Reactivity Trends in the Periodic Table**

**Reading Passage**

Read the passage carefully before answering the questions that follow.

The periodic table organizes elements based on their atomic number and properties, revealing patterns known as periodic trends. One of these trends is reactivity, which describes how likely an element is to participate in chemical reactions. Understanding these trends helps scientists predict how elements will behave in various chemical contexts.

Metals, such as sodium, calcium, and potassium, are highly reactive because they have a single valence electron that they readily lose in chemical reactions. This tendency is due to their large atomic size and weak nuclear energy, making it energetically favorable for them to react with other substances, especially nonmetals. Conversely, noble gases, like helium, neon, and argon, are very inert because they have full valence electron shells. This stable electronic configuration makes them unlikely to form bonds or participate in reactions under normal conditions.

Across periods in the periodic table, reactivity generally decreases as atomic number increases and valence electrons are held more tightly. In addition, atomic size influences reactivity; larger atoms tend to be more reactive if they have fewer valence electrons in their outermost shell. This electronic configuration directly impacts their ability to attract or give up electrons during chemical reactions. The periodic table also shows that reactivity trends are not uniform across all elements. For example, nonmetals like fluorine and oxygen are highly reactive, while transition metals show more varied reactivity patterns. These trends are crucial for understanding the chemical behavior of elements and predicting the outcomes of chemical reactions, especially in industrial settings, environmental science, and materials research.

**True or False: Reactivity in the Periodic Table**

Circle whether each statement is True or False.

- Metals are highly reactive because they have single valence electrons that are difficult to lose.
- Noble gases are very reactive because they have full valence electron shells.
- Reactivity increases across a period, which generally decreases the number of valence electrons.
- Metals and nonmetals show similar reactivity trends, such as being highly reactive.

# Periodic Trends Reactivity Lab Answer Key

**Heidi Andrade, Gregory J. Cizek**



## **Periodic Trends Reactivity Lab Answer Key:**

**Handbook of Formative Assessment** Heidi Andrade, Gregory J. Cizek, 2010-04-02 Formative assessment has recently become a focus of renewed research as state and federal policy makers realize that summative assessments have reached a point of diminishing returns as a tool for increasing student achievement. Consequently, supporters of large scale testing programs are now beginning to consider the potential of formative assessments to improve student achievement. The mission of this handbook is to comprehensively profile this burgeoning field of study. Written by leading international scholars and practitioners, each chapter includes a discussion of key issues that dominate formative assessment policy and practice today, as well as those that are likely to affect research and practice in the coming years. Key features include: Comprehensive nineteen chapters cover all aspects of formative assessment including classroom assessment, large scale applications, technological applications, applications for special needs students, K-12 and post-secondary applications, psychometric considerations, case studies, and discussion of alternative assessment formats such as portfolios and performance assessments. Integrative thoughtful attention is given to the integration of large scale and classroom assessments. Practical provides practical guidance on how to conduct formative assessments that generate credible information to guide instruction. Global provides perspectives from leading international scholars and practitioners whose expertise spans diverse settings, student populations, and educational systems. Accessible Style although grounded in the latest research, the book's style and tone has been carefully crafted to make it accessible to both the textbook and professional markets. It will also be a critical reference book for researchers in teacher preparation, educational administration, and educational policy studies.

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