

## Acids, Bases, Buffers and pH Data Tables

### Part A.

	Solution	pH Meter		pH paper		Universal Indicator	
		Reading	$[H_3O^+]$	color	pH value	color	pH value
	Control ( $H_2O$ )						
a.	0.1 M $NH_3$	11.00	$10^{-11} = 1^{-11}$	blue	12	purple	10
b.	0.1 M NaOH	12.85	$10^{-12.85} = 1.41 \times 10^{-13}$	blue	13	purple	10
c.	0.1 M $H_3BO_3$	6.5	$10^{-6.5} = 3.16 \times 10^{-7}$	orange yellow	4.5	yellow	6.5
d.	0.1 M HCl	1.40	$10^{-1.40} = .039$	orange	3	red/pink	3

### Part B.

	Solution	pH Meter		Universal Indicator	
		Reading	$[H_3O^+]$	color	pH value
	Control ( $H_2O$ )				
a.	0.1 M $Na_2CO_3$	11.42	$10^{-11.42} = 3.80 \times 10^{-12}$	purple	10
b.	0.1 M $(NH_4)_2SO_4$	6.58	$10^{-6.58} = 2.63 \times 10^{-7}$	yellow	6
c.	0.1 M NaCl	6.80	$10^{-6.80} = 1.58 \times 10^{-7}$	bright green	6.5

### Part C.

	Solution	pH Meter
		Reading
	Control ( $H_2O$ )	
a.	Acetic Acid	3.05
b.	Sodium Acetate	7.15
c.	Mixture of a & b	4.83

# Ph Properties Of Buffer Solutions Pre Lab Answers

**E Durkheim**



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