

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**
  - *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) = < 0.3)
    - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
  - *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
    - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
  - *Ionic Bond* ( $\Delta EN = > 1.7$ )
    - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
    - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
  - *Metallic Bond*
    - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- \* Must memorize/know the difference in electronegativity ranges for each type of bond
  - \*\* Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
  - See the table at the bottom of Packet pg.5
  - You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
  - Why → because all atoms want to be at their LOWEST POTENTIAL ENERGY; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
  - How → ATTRACTIVE and REPULSIVE forces
    - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
    - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
  - **Bond Length**
    - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
  - **Bond Energy**
    - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chapter 6 Study Guide Chemistry

**Huangqi Zhang**



## Chapter 6 Study Guide Chemistry:

**Hazmat Chemistry Study Guide (Second Edition)** Jill Meryl Levy, 2011      **ASAP Chemistry: A Quick-Review Study Guide for the AP Exam** The Princeton Review, 2019-02-12 Looking for sample exams practice questions and test taking strategies Check out our extended in depth AP chem prep guide Cracking the AP Chemistry Exam LIKE CLASS NOTES ONLY BETTER The Princeton Review s ASAP Chemistry is designed to help you zero in on just the information you need to know to successfully grapple with the AP test No questions no drills just review Advanced Placement exams require students to have a firm grasp of content you can t bluff or even logic your way to a 5 Like a set of class notes borrowed from the smartest student in your grade this book gives you exactly that No tricks or crazy stratagems no sample essays or practice sets Just the facts presented with lots of helpful visuals Inside ASAP Chemistry you ll find Essential concepts terms and functions for AP Chem all explained clearly concisely Diagrams charts and graphs for quick visual reference A three pass icon system designed to help you prioritize learning what you MUST SHOULD and COULD know in the time you have available Ask Yourself questions to help identify areas where you might need extra attention A resource that s perfect for last minute exam prep and for daily class work Topics covered in ASAP Chemistry include Atomic structure Covalent bonding intermolecular forces Thermochemistry Acids bases and more      *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH<sub>4</sub> from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement      **Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed** Robert K. Wismer, 1982      General Chemistry Study Guide Sixth Edition Darrell D. Ebbing, 1999      **Study Guide and Solutions Manual to Accompany Organic Chemistry, 11th Edition** T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2013-03-25 This is the study guide and solutions manual to accompany Organic Chemistry 11th Edition      **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines      **Study Guide for Chemistry by Steven S. Zumdahl** Martha B.

Barrett,1986      **Excel Science Study Guide, Years 9-10** Will Marchment,2004 The book contains coverage of five major topic areas in the NSW School Certificate test Energy Force and Motion Atoms Elements and Compounds Structure and Function of Living Things Earth and Space Ecosystems Resources and Technology a chapter on Investigations and Problem Solving in Science to help with practical skills revision questions and chapter tests to help you remember important information a glossary and summary in each section of the book diagrams and illustrations to help your understanding a section to help you prepare for the School Certificate test a sample School Certificate test paper with answers answers to all questions      **Study Guide and Solutions Manual for Fundamentals of Organic Chemistry** Susan McMurry,1994 Following a brief review of structure and bonding organic molecules and functional groups are presented as early as possible The text is organized primarily by functional group beginning with simple alkanes and moving toward more complex compounds Emphasis is placed on the fundamental mechanistic similarities of organic reactions McMurry's thorough revision continues to present the solid content necessary for this course without sacrifice of important subjects and pedagogical tools Text and reaction summaries full problem sets and outstanding artwork are just some of the features in the Third Edition usually found in a full year book McMurry's clear well written explanations remain a highlight of the book

*Study Guide [to Accompany] General Chemistry* Dixie Goss,2002      **Study Guide, Answer Book to Accompany Organic Chemistry** Alan S. Wingrove,Robert L. Caret,1981      **Essential Concepts of Chemistry Study Guide** James R. Braun,Sherman,1999      *Pharmacotherapy Principles and Practice Study Guide* Michael D. Katz,Kathryn R. Matthias,Marie A. Chisholm-Burns,2010-12-17 A case based companion study guide to Pharmacotherapy Principles and Practice 2e learn how to apply your knowledge to actual patient situations Pharmacotherapy Principles and Practice Study Guide uses 98 cases to help you learn how to apply pharmacotherapeutic concepts to specific patient situations Each case is presented in a consistent manner similar to what you would see in a clinical setting and focuses on one primary topic or problem Patients discussed in these cases will have drug therapy problems requiring identification and management For each case you will be asked to develop a Patient Database Drug Therapy Problem Worksheet and Pharmacotherapy Care Plan using the forms provided These forms are adapted from those originally developed by the American Society of Health System Pharmacists Clinical Skills program Each case includes Learning Objectives Patient Presentation Targeted Questions followed by a hint that refers you to pages in Pharmacotherapy Principles and Practice 2e where you can find the information to answer the question Follow up Global Perspective which highlights an issue related to the case that is important to countries outside of North America or involve different ethnic groups or races Case Summary      **Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition** G. Marc Loudon,Joseph G. Stowell,1995      **Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson** Norman E. Griswold,Henry Fuller Holtzclaw,1988      Study Guide and Solutions Manual for Organic Chemistry Susan McMurry,1992 John McMurry's best

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100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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