

Calculate the Lattice Enthalpy of  $\text{CaCl}_2$  crystal

(a) Using Born Hyber Cycle

(b) Using Kapunstisnskii Equation

$$\Delta H_L = -1.21 \frac{nZ_A Z_B}{d} \left( 1 - \frac{d^*}{d} \right) \quad \text{MJ} \cdot \text{\AA} \cdot \text{mol}^{-1}$$

Ionic radii:  $\text{Ca}^{+2}$ : 1.0 \AA ;  $\text{Cl}^-$ : 1.67 \AA

Enthalpy of formation for  $\text{CaCl}_2 = -795 \text{ kJ}$

Sublimation of Ca : 325 kJ

Dissociation enthalpy of  $\text{Cl}_2$ : 244 kJ

Electron Affinity for Cl: 355 kJ

Total Ionization enthalpy for  $\text{Ca(g)} \rightarrow \text{Ca}^{+2}(\text{g})$  : 1735 kJ

$d^* = 0.345 \text{ \AA}$

# Enthalpy Of Solution Cacl

**Allan Myerson**



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The NBS Tables of Chemical Thermodynamic Properties Donald D. Wagman, 1982

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