

### 7B-3 : Entropy Change of an Isobaric Process

6 pts

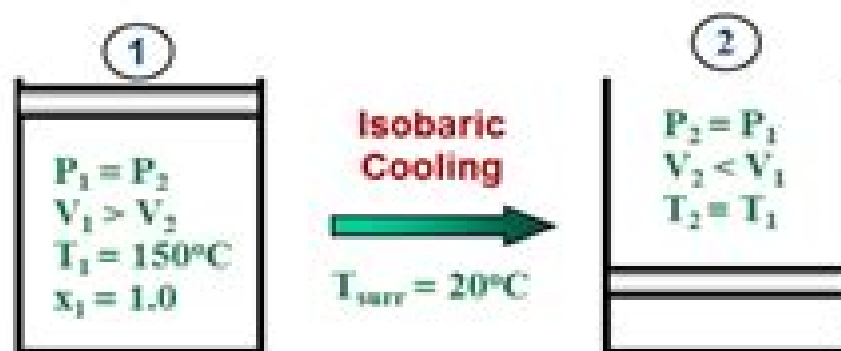
A piston-and-cylinder device with a free-floating piston contains 2.6 kg of saturated water vapor at 150°C. The water loses heat to the surroundings until the cylinder contains saturated liquid water.

The surroundings are at 20°C. Calculate... a.)  $\Delta S_{\text{water}}$ , b.)  $\Delta S_{\text{surroundings}}$ , c.)  $\Delta S_{\text{universe}}$

**Read :** Calculating  $\Delta S$  for the water in the cylinder is straightforward. The key to calculating  $\Delta S_{\text{surr}}$  is the fact that the surroundings behave as a thermal reservoir. The temperature of the surroundings does not change and there are no irreversibilities in the surroundings that are associated with the process. The key to calculating  $\Delta S_{\text{univ}}$  is the fact that the universe is made up of the combination of the system and the surroundings. Consequently,  $\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$ .

<b>Given:</b>	$m$	2.60	kg	<b>Find:</b>	$\Delta S_{\text{sys}}$	???	kJ/K
	$T_1$	150	°C		$\Delta S_{\text{surr}}$	???	kJ/K
	$x_1$	1	kg vap/kg		$\Delta S_{\text{univ}}$	???	kJ/K
	$T_{\text{surr}}$	20	°C				
	$P_1 = P_2$						
	$x_2$	0	kg vap/kg				

**Diagram:**



- Assumptions:**
1. Changes in kinetic and potential energies are negligible.
  2. Boundary work is the only form of work that crosses the system boundary.
  3. The surroundings behave as a thermal reservoir.

# Entropy Problems And Solutions

**DP Hallahan**



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