



Student Exploration: Balancing Chemical Equations

(This document is intended for printing.)

Essential Question: How do we balance chemical equations? What are the coefficients in a balanced chemical equation? How do we use a balanced chemical equation to predict the products of a chemical reaction?

What You Will Do: You will use the law of conservation of mass to balance chemical equations.

1. You will use the law of conservation of mass to balance chemical equations. You will use the law of conservation of mass to predict the products of a chemical reaction.
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What You Will Do

You will use the law of conservation of mass to balance chemical equations. You will use the law of conservation of mass to predict the products of a chemical reaction.

Essential Question: How do we balance chemical equations? What are the coefficients in a balanced chemical equation? How do we use a balanced chemical equation to predict the products of a chemical reaction?

What You Will Do: You will use the law of conservation of mass to balance chemical equations.

Example: $2H_2 + O_2 \rightarrow 2H_2O$

1. Look at the unbalanced equation above. It is not balanced because the number of atoms of each element is not the same on both sides of the equation. The coefficients tell you how many molecules of each substance are involved in the reaction. The number of atoms of each element is the same on both sides of the equation.

Hydrogen: 4 atoms Oxygen: 2 atoms Water: 4 atoms

2. It is a balanced equation. The number of atoms of each element is the same on both sides of the equation. The number of atoms of each element is the same on both sides of the equation.

Hydrogen: 4 atoms Oxygen: 2 atoms Water: 4 atoms
Hydrogen: 4 atoms Oxygen: 2 atoms Water: 4 atoms

How do we balance a chemical equation? First, we identify the elements.

What You Will Do

Example: $2H_2 + O_2 \rightarrow 2H_2O$

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What You Will Do



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Christian G. Meyer



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