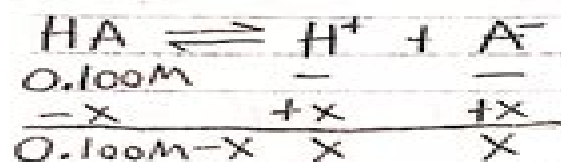


• Calculate K_a for the following acids:

1) 100.0 mL of 0.100 M HA with a pH of 3.40

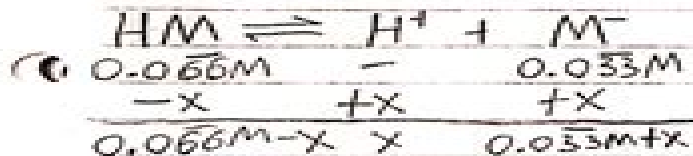


$$\begin{aligned}
 \text{pH} &= 3.40 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 3.40 &= -\log(x) \rightarrow x = 3.98 \times 10^{-4}
 \end{aligned}$$

$$K_a = \left(\frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left(\frac{(x)^2}{(0.100\text{M}-x)} \right)$$

$$K_a = \left(\frac{(3.98 \times 10^{-4})^2}{(0.100\text{M} - 3.98 \times 10^{-4})} \right) \rightarrow \boxed{K_a = 1.59 \times 10^{-6}}$$

2) 100.0 mL of 0.100 M HM is mixed with 50.0 mL of 0.100 M NaM and has a pH of 5.50



$$\begin{aligned}
 \text{pH} &= 5.50 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 5.50 &= -\log(x) \rightarrow x = 3.16 \times 10^{-6}
 \end{aligned}$$

Dilution problem due to HM and NaM

$$\begin{aligned}
 (C_1)(V_1) &= (C_2)(V_2) \\
 (0.100\text{M HM})(100.0\text{mL}) &= (C_2)(150.0\text{mL}) \\
 (C_2) &= 0.066\text{M HM}
 \end{aligned}$$

Dilution problem due to NaM and HM

$$\begin{aligned}
 (C_1)(V_1) &= (C_2)(V_2) \\
 (0.100\text{M NaM})(50.0\text{mL}) &= (C_2)(150.0\text{mL}) \\
 (C_2) &= 0.033\text{M NaM}
 \end{aligned}$$

$$K_a = \left(\frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left(\frac{(x)(0.033\text{M}+x)}{(0.066\text{M}-x)} \right)$$

$$K_a = \left(\frac{(3.16 \times 10^{-6})(0.033\text{M} + 3.16 \times 10^{-6})}{(0.066\text{M} - 3.16 \times 10^{-6})} \right) \rightarrow K_a = 1.58 \times 10^{-6}$$

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Terry C. Jones



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1. Understanding the eBook College Chemistry Problems And Solutions
 - The Rise of Digital Reading College Chemistry Problems And Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying College Chemistry Problems And Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an College Chemistry Problems And Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from College Chemistry Problems And Solutions
 - Personalized Recommendations
 - College Chemistry Problems And Solutions User Reviews and Ratings
 - College Chemistry Problems And Solutions and Bestseller Lists
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 - Highlighting and Note-Taking College Chemistry Problems And Solutions
 - Interactive Elements College Chemistry Problems And Solutions
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- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers College Chemistry Problems And Solutions
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 - Creating a Diverse Reading Collection College Chemistry Problems And Solutions
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine College Chemistry Problems And Solutions
- Setting Reading Goals College Chemistry Problems And Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of College Chemistry Problems And Solutions
- Fact-Checking eBook Content of College Chemistry Problems And Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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