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Introduction

Molecules have shape! The structure and shape of a molecule influences its physical properties and affects its chemical behavior as well. Lewis structures and VSEPR theory offer useful models for visualizing the structures of covalent compounds.

Concepts

- Valence electrons
- Lewis Dot structures
- Covalent bonding
- VSEPR theory

Background

Covalent bonds are defined as the net attractive forces between nonmetal atoms that share one, two, or three pairs of electrons. In general, only the *valence electrons*, those in the highest energy levels that are the farthest away from the nucleus, are available for bonding. The number of valence electrons influences the number of bonds that an atom will form. The periodic table offers a convenient shortcut for determining the number of valence electrons in an atom. Remember that the position of an element in the modern periodic table reflects its electrons configuration. When the elements are arranged in columns from IA to VIIIA, the number of valence electrons for an element is equal to its group number. Thus, potassium in Group IA has one valence electron, carbon in Group IVA has four valence electrons, and chlorine in Group VIIA has seven valence electrons. In 1916, G.N. Lewis, an American chemist, proposed arranging dots around the symbols of the elements to represent the valence electrons. *Lewis dot symbols* remain the most popular way to picture the valence electrons that are available for bonding.



Lewis dot structures for molecules build on the Lewis dot symbols of the elements to show the bonding arrangement of atoms in a molecule and the distribution of all valence electrons. The Lewis structure of a molecule thus shows all of the atoms and how they are connected. A single covalent bond between two atoms is represented using two dots. Sometimes atoms share more than one pair of electrons between them. Four dots are used to represent a double

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Ian Pickup



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