

Calculate the Lattice Enthalpy of  $\text{CaCl}_2$  crystal

(a) Using Born Hyber Cycle

(b) Using Kapunstisnskii Equation

$$\Delta H_L = -1.21 \frac{nZ_A Z_B}{d} \left( 1 - \frac{d^*}{d} \right) \quad \text{MJ} \cdot \text{\AA} \cdot \text{mol}^{-1}$$

Ionic radii:  $\text{Ca}^{+2}$ : 1.0 \AA ;  $\text{Cl}^-$ : 1.67 \AA

Enthalpy of formation for  $\text{CaCl}_2 = -795 \text{ kJ}$

Sublimation of Ca : 325 kJ

Dissociation enthalpy of  $\text{Cl}_2$ : 244 kJ

Electron Affinity for Cl: 355 kJ

Total Ionization enthalpy for  $\text{Ca}(\text{g}) \rightarrow \text{Ca}^{+2}(\text{g})$  : 1735 kJ

$d^* = 0.345 \text{ \AA}$

# Enthalpy Of Solution Cacl

**Victor M. Corman**



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### **Table of Contents Enthalpy Of Solution Cacl**

1. Understanding the eBook Enthalpy Of Solution Cacl
  - The Rise of Digital Reading Enthalpy Of Solution Cacl
  - Advantages of eBooks Over Traditional Books
2. Identifying Enthalpy Of Solution Cacl
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Enthalpy Of Solution Cacl
  - User-Friendly Interface
4. Exploring eBook Recommendations from Enthalpy Of Solution Cacl
  - Personalized Recommendations

- Enthalpy Of Solution Cacl User Reviews and Ratings
- Enthalpy Of Solution Cacl and Bestseller Lists
- 5. Accessing Enthalpy Of Solution Cacl Free and Paid eBooks
  - Enthalpy Of Solution Cacl Public Domain eBooks
  - Enthalpy Of Solution Cacl eBook Subscription Services
  - Enthalpy Of Solution Cacl Budget-Friendly Options
- 6. Navigating Enthalpy Of Solution Cacl eBook Formats
  - ePub, PDF, MOBI, and More
  - Enthalpy Of Solution Cacl Compatibility with Devices
  - Enthalpy Of Solution Cacl Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Enthalpy Of Solution Cacl
  - Highlighting and Note-Taking Enthalpy Of Solution Cacl
  - Interactive Elements Enthalpy Of Solution Cacl
- 8. Staying Engaged with Enthalpy Of Solution Cacl
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Enthalpy Of Solution Cacl
- 9. Balancing eBooks and Physical Books Enthalpy Of Solution Cacl
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Enthalpy Of Solution Cacl
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Enthalpy Of Solution Cacl
  - Setting Reading Goals Enthalpy Of Solution Cacl
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Enthalpy Of Solution Cacl
  - Fact-Checking eBook Content of Enthalpy Of Solution Cacl

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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