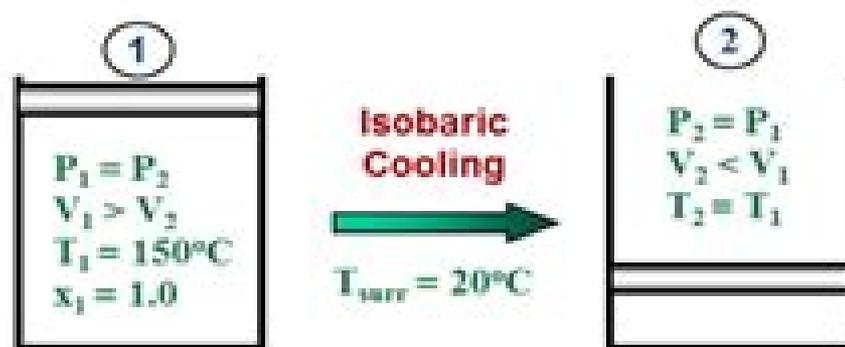## 7B-3 : Entropy Change of an Isobaric Process

6 pts

A piston-and-cylinder device with a free-floating piston contains 2.6 kg of saturated water vapor at 150°C. The water loses heat to the surroundings until the cylinder contains saturated liquid water.

The surroundings are at 20°C. Calculate... a.) ΔS_{water}, b.) $\Delta S_{surroundings}$, c.) $\Delta S_{universe}$

Read : Calculating ΔS for the water in the cylinder is straightforward. The key to calculating ΔS_{surr} is the fact that the surroundings behave as a thermal reservoir. The temperature of the surroundings does not change and there are no irreversibilities in the surroundings that are associated with the process. The key to calculating ΔS_{univ} is the fact that the universe is made up of the combination of the system and the surroundings. Consequently, $\Delta S_{univ} = \Delta S_{sys} + \Delta S_{surr}$.

Given:

m	2.60	kg
T_1	150	°C
x_1	1	kg vap/kg
T_{surr}	20	°C
$P_1 = P_2$		
x_2	0	kg vap/kg

Find:

ΔS_{sys}	???	kJ/K
ΔS_{surr}	???	kJ/K
ΔS_{univ}	???	kJ/K

Diagram:

Assumptions:

1 - Changes in kinetic and potential energies are negligible.
2 - Boundary work is the only form of work that crosses the system boundary.
3 - The surroundings behave as a thermal reservoir.

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