

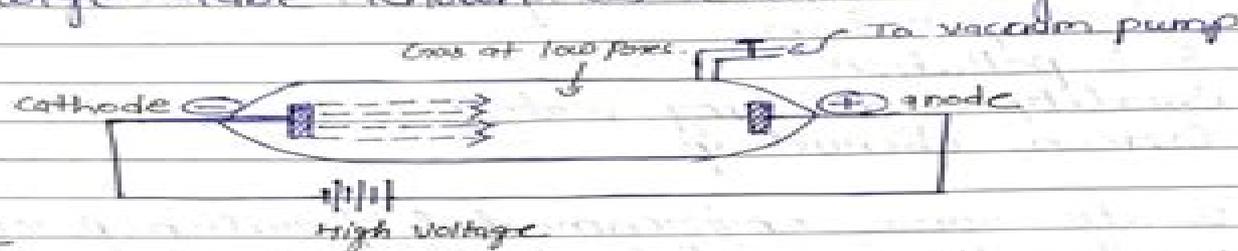
## CHAPTER - 2

### STRUCTURE OF ATOM

#### Discovery of electron (Cathode Rays)

Cathode rays  $\rightarrow$  William Crookes (1879)  
Electron ( $e^-$ )  $\rightarrow$  J. J. Thomson (-ve charge particle)

-ve charged rays which move from cathode to anode in discharge tube known as cathode rays.



#### properties

- 1 Cathode rays travel in a straight line (shadow)
- 2 Cathode rays carry -ve charge (electrostatic field deflection towards anode)
- 3 Cathode rays are deflected by magnetic field (rotate paddle wheel)
- 4 They are made up of material particle
- 5 They produce heating effect & possess penetrating effect
- 6 They produce green fluorescence & X-rays (when they strike metal)
- 7 They also affect photographic plate.

#### Discovery of proton (Anode / Canal Rays)

Anode rays  $\rightarrow$  Goldstein (1886)  
Proton ( $H^+$ )  $\rightarrow$  Goldstein & named by Rutherford (+ve charge particle)

+ve charged particle which move opposite the cathode in discharge tube known as anode rays.

# Class 11th Notes Chemistry Chapter Wise

**Lei Shi**



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### **Table of Contents Class 11th Notes Chemistry Chapter Wise**

1. Understanding the eBook Class 11th Notes Chemistry Chapter Wise
  - The Rise of Digital Reading Class 11th Notes Chemistry Chapter Wise
  - Advantages of eBooks Over Traditional Books
2. Identifying Class 11th Notes Chemistry Chapter Wise
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Class 11th Notes Chemistry Chapter Wise
  - User-Friendly Interface
4. Exploring eBook Recommendations from Class 11th Notes Chemistry Chapter Wise

- Personalized Recommendations
  - Class 11th Notes Chemistry Chapter Wise User Reviews and Ratings
  - Class 11th Notes Chemistry Chapter Wise and Bestseller Lists
5. Accessing Class 11th Notes Chemistry Chapter Wise Free and Paid eBooks
    - Class 11th Notes Chemistry Chapter Wise Public Domain eBooks
    - Class 11th Notes Chemistry Chapter Wise eBook Subscription Services
    - Class 11th Notes Chemistry Chapter Wise Budget-Friendly Options
  6. Navigating Class 11th Notes Chemistry Chapter Wise eBook Formats
    - ePub, PDF, MOBI, and More
    - Class 11th Notes Chemistry Chapter Wise Compatibility with Devices
    - Class 11th Notes Chemistry Chapter Wise Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Class 11th Notes Chemistry Chapter Wise
    - Highlighting and Note-Taking Class 11th Notes Chemistry Chapter Wise
    - Interactive Elements Class 11th Notes Chemistry Chapter Wise
  8. Staying Engaged with Class 11th Notes Chemistry Chapter Wise
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Class 11th Notes Chemistry Chapter Wise
  9. Balancing eBooks and Physical Books Class 11th Notes Chemistry Chapter Wise
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Class 11th Notes Chemistry Chapter Wise
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Class 11th Notes Chemistry Chapter Wise
    - Setting Reading Goals Class 11th Notes Chemistry Chapter Wise
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Class 11th Notes Chemistry Chapter Wise

- Fact-Checking eBook Content of Class 11th Notes Chemistry Chapter Wise
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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