

THE MOLE: LAB ACTIVITY

Station 1:

A) Determine the mass of one drop of water by measuring the mass of 50 drops of water.

Mass of dish (or beaker)	= _____ g
Mass of dish (or beaker) + 50 drops water	= _____ g
Mass of 50 drops water	= _____ g
Mass of 1 drop of water	= _____ g

B) How many molecules are there in one drop of water?

Remember 1 mole of a substance is 6.02×10^{23} particles (atoms or molecules).
Therefore, 6.02×10^{23} molecules of water will weigh 18.02 g/mol

1. Find # of moles of water for 1 drop $n = (\text{mass})/(\text{Molar Mass})$
2. # of molecules in one drop of water $= n_{\text{water}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$
 $= ? \text{ molecules}$

Station 2

Calculate the number of copper atoms in a penny.

1. Find # of moles of Cu $n = (\text{mass})/(\text{Molar Mass})$
2. # of Cu atoms in a penny $= n_{\text{Cu}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$
 $= ? \text{ atoms}$

Station 3

A. Calculate the number of moles in one spoon of sugar. Assume that sugar is made up entirely of sucrose molecules ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$).

$$M_{\text{sucrose}} = 342.34 \text{ g/mol}$$

$$n_{\text{sugar}} = \frac{M_{\text{sucrose}}}{M_{\text{sucrose}}}$$
$$= \underline{\hspace{2cm}} \text{ mol}$$

B. Calculate the number of molecules of sugar in one spoon of sugar.

$$\begin{aligned} \# \text{ of molecules sucrose} &= n_{\text{sucrose}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \underline{\hspace{2cm}} \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \\ &= ? \text{ molecules} \end{aligned}$$

Mole Lab Chemistry I Acc Answers

J Rink



Mole Lab Chemistry I Acc Answers:

Berkeley Lab Research Review, 1997 **Cooperative Chemistry Lab Manual** Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

Study Guide for Hein and Arena's Foundations of College Chemistry Peter Scott, Rachael Enriquez Porter, 1999-04-18

The Success Manual for General Chemistry Elizabeth Kean, Catherine Middlecamp, 1986 **Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors **How to Survive (and Even Excel In) General Chemistry** Elizabeth Kean, Catherine Middlecamp, 1994 A different kind of book about chemistry which teaches readers the process of learning chemistry not the topic itself Proving a valuable supplement to any introductory text this guide offers inside information to help make chemistry less stressful even enjoyable Includes exercises and sections for self assessment Student Study Guide to accompany Foundations of College Chemistry, 11th Edition Morris Hein, Susan Arena, 2003-09-03 Work more effectively and gauge your progress along the way This Student Study Guide that accompanies Hein's Foundations of College Chemistry 11th Edition contains a review and self evaluation section challenge problems and answers and solutions for all exercises in the study guide Foundations of College

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