

Stoichiometric Calculations: Identify a compound using gravimetric analysis

Question 1: Atoms and molecules are way too small to be seen, even with a microscope, so how can we count how many there are?

- ✓ You count them by mass
- You count them using an electron microscope
- You can't count atoms and molecules
- You count them by volume

Question 2: In which of the pictures do you find the most particles? By particles I mean, sand grains, water molecules and iron atoms.

- ✓ There are roughly the same amount of particles on each picture
- In the iron bolt! It is quite heavy, and iron atoms are very small
- In the water glass! A shot of water isn't very large, but water molecules are super tiny
- In the Sahara desert! The desert is huge, and sand grains are small

Question 3: Why is it important to use demineralized water?

- ✓ Because demineralized water doesn't contain minerals
- Because demineralized water has a higher boiling point than regular tap-water
- Because demineralized water poses less of a health hazard in chemical laboratories
- Because demineralized water has a lower density than regular tap-water

Question 4: In order for the gravimetric analysis to be valid, we want all of the chloride ions to precipitate. How can we ensure this?

- ✓ By ensuring silver ions are in excess
- By adding the same volume of reactants
- By having roughly the same amount of each reactant
- By ensuring chloride ions are in excess

Question 5: Look at the reaction schemes in your journal or click the "View image" button below. How does the number of moles of AgNO_3 silver nitrate correlate to the number of moles of the alkaline earth metal chloride?

- ✓ There should be twice as many moles of silver nitrate AgNO_3 as unknown chloride XCl_2
- We cannot tell before we have performed the reaction
- There should be half as many moles of silver nitrate AgNO_3 as unknown chloride XCl_2
- There should be the same number of moles of each

Question 6: How many moles of AgNO_3 silver nitrate should we use to be sure that we have excess, no matter which of the three compounds it is?

- ✓ 0.100 mol
- 0.075 mol
- 0.050 mol
- 0.025 mol

Question 7: Look at your lab journal: Can you see from the reaction stoichiometry how the number of moles of AgCl silver chloride formed correlates to the number of moles of AgNO_3 silver nitrate consumed?

- ✓ The number of moles are the same
- It differs in the three different reactions
- There are half as many moles of AgCl formed
- There are twice as many moles of AgCl formed

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J Ma



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