

KEY

1. Complete and classify the following partial equations

- a) $1 \text{ Fe} + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe SO}_4 + 3 \text{ H}_2$ (single repl)
- b) $\text{NaI} + \text{Cl}_2 \rightarrow$
- c) $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$ (single repl)
- d) $\text{Al(OH)}_3 + \text{HCl} \rightarrow$
- e) $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$ $\text{4 Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$ (synthesis)
- f) $\text{Al}_2\text{O}_3 \rightarrow$
- g) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ (synth)
- h) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- i) $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$ (s.r.)
- j) $\text{Na} + \text{H}_2\text{O} \rightarrow$
- k) $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ (acid base)
- l) $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
- m) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2 \text{ CO}_2 + 3 \text{ H}_2\text{O}$ (combustion)
- n) $\text{AlCl}_3 \rightarrow$
- o) $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$ (acid base)
- p) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- q) $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$ (acid base)
- r) $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
- s) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ (combustion)
- t) $\text{H}_2\text{O} \rightarrow$
- u) $2 \text{ HNO}_3 + \text{CuCO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 1 \text{ CO}_2$ (s.r.)
- v) $\text{CH}_4 + \text{O}_2 \rightarrow$
- w) $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$ (s.r.)
- x) $\text{KI} + \text{Br}_2 \rightarrow$
- y) $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$ (acid base)
- z) $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$



3 each

2 each

3

4

1/2/1/1/1

1/2/1

1/2/1/1

5

2. Given: $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$ calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$)

254g

4. If 1.40 moles of butane (C_4H_{10}) is burned

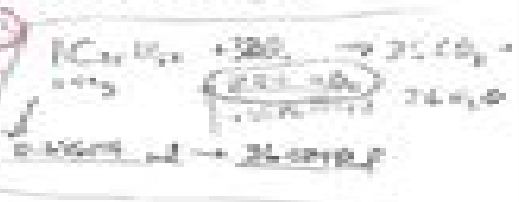
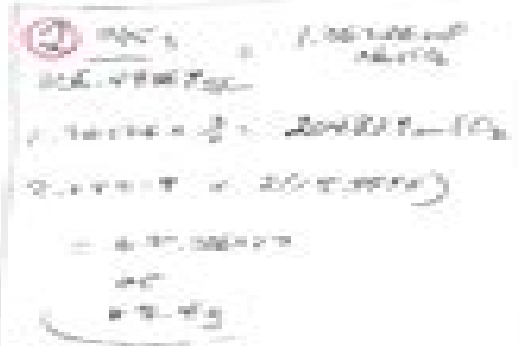
- a) calculate the number of moles of oxygen gas required (9.8 moles)
- b) calculate the number of moles of carbon dioxide produced (10.8 moles CO_2)
- c) calculate the mass of water produced (10.8 moles H_2O)

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate

- a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. (27.4g Al)
- b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. (192 g Zn)

6. Given: $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$

- a) How many grams of Cu are required to react with 3.00 g of silver nitrate? (0.544g Cu)
- b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? (1.70g Ag)



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1. Understanding the eBook Nelson Chemistry 30 Answer Key
 - The Rise of Digital Reading Nelson Chemistry 30 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 30 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 30 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 30 Answer Key
 - Personalized Recommendations
 - Nelson Chemistry 30 Answer Key User Reviews and Ratings
 - Nelson Chemistry 30 Answer Key and Bestseller Lists

5. Accessing Nelson Chemistry 30 Answer Key Free and Paid eBooks
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 - Nelson Chemistry 30 Answer Key Budget-Friendly Options
6. Navigating Nelson Chemistry 30 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Nelson Chemistry 30 Answer Key Compatibility with Devices
 - Nelson Chemistry 30 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Chemistry 30 Answer Key
 - Highlighting and Note-Taking Nelson Chemistry 30 Answer Key
 - Interactive Elements Nelson Chemistry 30 Answer Key
8. Staying Engaged with Nelson Chemistry 30 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 30 Answer Key
9. Balancing eBooks and Physical Books Nelson Chemistry 30 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 30 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry 30 Answer Key
 - Setting Reading Goals Nelson Chemistry 30 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry 30 Answer Key
 - Fact-Checking eBook Content of Nelson Chemistry 30 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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