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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} + \text{CO} \rightarrow \text{CO}_2 + \text{Fe}$
oxidation $\text{CO} \rightarrow \text{CO}_2 + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ag} + \text{Ni}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$ $\text{O}_5 + 5\text{C} \rightarrow \text{O}_2 + 5\text{CO}_2$

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reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^2 oxidized reducing agent O_0 to O^2 reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^1 oxidized red ag S^0 to S^2 reduced ox ag $3\text{Cs} + \text{Br}_2$

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