

Stoichiometry III – Theoretical Yield

Teacher

If we know what chemical reaction is occurring, we can use mole ratios to determine the limiting and excess reactants. We can do this by using the starting masses of each reactant to see how much of a product could be produced if all of the reactant was consumed. We then compare the number of moles produced by each reactant, and the smallest value is the theoretical yield – the amount of a product produced if all of the limiting reactant is consumed.

Practice Problems



a) Determine the limiting and excess reagents if you started with 14.8 g of C_3H_8 and 33.44 g of O_2 .

$$\frac{14.8 \text{ g } C_3H_8}{44.09 \text{ g } C_3H_8} \times \frac{1 \text{ mol } C_3H_8}{1 \text{ mol } C_3H_8} \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } C_3H_8} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 24.2 \text{ g } H_2O$$

$$\frac{33.44 \text{ g } O_2}{32.00 \text{ g } O_2} \times \frac{1 \text{ mol } O_2}{5 \text{ mol } O_2} \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } O_2} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 15.1 \text{ g } H_2O$$

Limiting: O_2
Excess: C_3H_8

b) What is the theoretical yield of H_2O . $15.1 \text{ g } H_2O$



a) Determine the limiting and excess reagents if 8.1 g of $Al_2(SO_4)_3$ is reacted with 8.1 g of $NaOH$

$$\frac{8.1 \text{ g } Al_2(SO_4)_3}{294.14 \text{ g } Al_2(SO_4)_3} \times \frac{1 \text{ mol } Al_2(SO_4)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{2 \text{ mol } Al(OH)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 4.3 \text{ g } Al(OH)_3$$

$$\frac{8.1 \text{ g } NaOH}{40.00 \text{ g } NaOH} \times \frac{1 \text{ mol } NaOH}{6 \text{ mol } NaOH} \times \frac{2 \text{ mol } Al(OH)_3}{1 \text{ mol } NaOH} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 5.3 \text{ g } Al(OH)_3$$

Limiting: $Al_2(SO_4)_3$
Excess: $NaOH$

b) What is the theoretical yield of aluminum hydroxide. $4.3 \text{ g } Al(OH)_3$



a) Determine the limiting and excess reagents if 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe .

$$\frac{25.4 \text{ g } Al_2O_3}{101.96 \text{ g } Al_2O_3} \times \frac{1 \text{ mol } Al_2O_3}{4 \text{ mol } Al_2O_3} \times \frac{3 \text{ mol } Fe_3O_4}{1 \text{ mol } Al_2O_3} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 43.3 \text{ g } Fe_3O_4$$

$$\frac{10.2 \text{ g } Fe}{55.85 \text{ g } Fe} \times \frac{1 \text{ mol } Fe}{9 \text{ mol } Fe} \times \frac{3 \text{ mol } Fe_3O_4}{1 \text{ mol } Fe} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 14.1 \text{ g } Fe_3O_4$$

Limiting: Fe
Excess: Al_2O_3

b) What is the theoretical yield of iron(II,III) oxide (aka magnetite).

$$14.1 \text{ g } Fe_3O_4$$

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