

## Chemistry Chapter 8 Test

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_ 1. What type of bond would you expect in a compound of oxygen and sodium?  
 a. ionic  
 b. metallic  
 c. covalent  
 d. polar covalent
- \_\_\_\_ 2. The symbols  $\delta^-$  and  $\delta^+$  indicate which of the following?  
 a. electronegativity  
 b. electrostatic force  
 c. charge  
 d. partial charge
- \_\_\_\_ 3. Which of the following represents a nonpolar covalent molecule?  
 a. NaCl  
 b. Br<sub>2</sub>  
 c. HCN  
 d. Au
- \_\_\_\_ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called \_\_\_\_\_.  
 a. electron affinity  
 b. polarity  
 c. electronegativity  
 d. ionization energy
- \_\_\_\_ 5. The polarity of a bond depends on the difference in the atoms' \_\_\_\_\_.  
 a. electron affinities  
 b. electronegativities  
 c. atomic radii  
 d. ionization energies
- \_\_\_\_ 6. What type of bond would you expect to form between nitrogen and potassium?  
 a. ionic  
 b. covalent  
 c. polar covalent  
 d. metallic
- \_\_\_\_ \* High electronegativity differences are found in \_\_\_\_\_ bonds.  
 a. metallic  
 b. ionic metal - nonmetal  
 c. covalent  
 d. Pauling
- \_\_\_\_ \* To achieve an octet when combining with potassium, would an atom of a Group 16 element be expected to lose, gain, or share electrons?  
 a. lose  
 b. gain  
 c. share  
 d. This cannot be determined from the given information.
- \_\_\_\_ 9. Which of the following does not exist as a diatomic molecule?  
 a. argon  
 b. hydrogen  
 c. chlorine  
 d. oxygen  
 H N O F Cl Br I  
 Gas: H<sub>2</sub> N<sub>2</sub> O<sub>2</sub> F<sub>2</sub> Cl<sub>2</sub> Br<sub>2</sub> I<sub>2</sub>
- \_\_\_\_ 10. Which of these is not a covalent molecule?  
 a. F<sub>2</sub>  
 b. CO  
 c. LiF  
 d. NH<sub>3</sub>
- \_\_\_\_ 11. A dash in a Lewis structure represents \_\_\_\_\_.  
 a. one electron  
 b. two electrons  
 c. four electrons  
 d. unpaired electrons
- \_\_\_\_ 12. The central atom in a Lewis structure is usually the atom \_\_\_\_\_.

# Chapter 5 Chemistry Test Answers

**Thomas Evangelist**



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## **Table of Contents Chapter 5 Chemistry Test Answers**

1. Understanding the eBook Chapter 5 Chemistry Test Answers
  - The Rise of Digital Reading Chapter 5 Chemistry Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 5 Chemistry Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 5 Chemistry Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 5 Chemistry Test Answers
  - Personalized Recommendations
  - Chapter 5 Chemistry Test Answers User Reviews and Ratings
  - Chapter 5 Chemistry Test Answers and Bestseller Lists
5. Accessing Chapter 5 Chemistry Test Answers Free and Paid eBooks
  - Chapter 5 Chemistry Test Answers Public Domain eBooks
  - Chapter 5 Chemistry Test Answers eBook Subscription Services

- Chapter 5 Chemistry Test Answers Budget-Friendly Options
- 6. Navigating Chapter 5 Chemistry Test Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 5 Chemistry Test Answers Compatibility with Devices
  - Chapter 5 Chemistry Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 5 Chemistry Test Answers
  - Highlighting and Note-Taking Chapter 5 Chemistry Test Answers
  - Interactive Elements Chapter 5 Chemistry Test Answers
- 8. Staying Engaged with Chapter 5 Chemistry Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 5 Chemistry Test Answers
- 9. Balancing eBooks and Physical Books Chapter 5 Chemistry Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 5 Chemistry Test Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 5 Chemistry Test Answers
  - Setting Reading Goals Chapter 5 Chemistry Test Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 5 Chemistry Test Answers
  - Fact-Checking eBook Content of Chapter 5 Chemistry Test Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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