

The Periodic Law

SECOND PERIODIC

NOTE: All answers should include the following questions in the space provided.

- Classify the central atoms in the following compounds as sp , sp^2 , sp^3 , or sp^3d hybridized.

CO_2 _____	a. Why is the central carbon sp hybridized?
CO_3^{2-} _____	b. Why is the central carbon sp^2 hybridized?
CH_4 _____	c. Is the chlorine in question (a) or question (b) hybridized? If so, how is it hybridized and by what orbital?
- Classify the following elements as sp , sp^2 , sp^3 , or sp^3d hybridized.

B _____	a. For which hybrid orbitals does the hybridization apply?
Si _____	b. Is it an sp , sp^2 , or sp^3 hybridized? How is it hybridized and by what orbital?
Te _____	c. The group number of the element is $6A$. How is it hybridized?
- | | |
|----------|---|
| B _____ | a. Is the hybridization sp , sp^2 , sp^3 , or sp^3d hybridized? |
| Si _____ | b. Is the hybridization sp , sp^2 , sp^3 , or sp^3d hybridized? How is it hybridized? |
- | | |
|----------|--|
| B _____ | a. Does the central atom hybridize? How is it hybridized? |
| Si _____ | b. Does the central atom hybridize? How is it hybridized? |
| Te _____ | c. Does the hybridization of the central atom hybridize? How is it hybridized? |
- Using only hybrid orbitals, write the hybrid orbitals for the following elements in the following compounds.

Be _____	a. Why does the hybridization sp hybridize?
Si _____	b. Why does the hybridization sp^3 hybridize?
Te _____	c. Why does the hybridization sp^3d hybridize?
Be _____	d. Why does the hybridization sp^3d hybridize?
Si _____	e. Why does the hybridization sp^3d hybridize?
Te _____	f. Why does the hybridization sp^3d hybridize?

Chapter 5 Review The Periodic Law Mixed Answer Key

M Tight



Chapter 5 Review The Periodic Law Mixed Answer Key:

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