

8

Acids, bases and salts

Key objectives

By the end of this section, you should be able to:

- describe the characteristic reactions of acids
 - state that bases are oxides or hydroxides of metals
 - state that alkalis are soluble bases
 - describe the characteristic reactions of bases
 - state that aqueous solutions of acids contain H^+ ions and aqueous solutions of alkalis contain OH^- ions
 - describe how to use pH, as measured with universal indicator paper, to compare hydrogen ion concentration, neutrality, relative acidity and alkalinity
 - describe the neutralisation reaction
- define acids and bases in terms of proton transfer
 - explain the difference between weak and strong acids in terms of dissociation
- classify oxides as acidic, basic or amphoteric, with examples
 - describe the general solubility rules for salts
 - describe the preparation of soluble salts by the reaction of an acid with:
 - an alkali
 - excess metal, insoluble base or carbonate
- describe the preparation of insoluble salts by precipitation
- define a hydrated substance and an anhydrous substance
- define the term water of crystallisation in crystals, including $CuSO_4 \cdot 5H_2O$ and $CoCl_2 \cdot 6H_2O$

Key terms

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Term	Definition
Acid	A substance which dissolves in water, producing $H^+(aq)$ ions as the only positive ion. A proton (H^+) donor.
Alkali	A soluble base which produces $OH^-(aq)$ ions in water.
Anhydrous salt	A salt which has had its water of crystallisation removed.
Base	A substance which neutralises an acid, producing a salt and water as the only products. Bases are oxides or hydroxides of metals. (Ammonia is also a base.) A proton (H^+) acceptor.
Indicator	A substance that shows whether a substance is acidic or alkaline by changing colour.
Neutralisation	The process in which an acid reacts with a base to form water.
pH scale	A scale running from 0 to 14 used to express the acidity or alkalinity of a substance.
Saturated solution	A solution containing the maximum amount of dissolved solute in the solvent at a given temperature.
Water of crystallisation	Water incorporated into the structure of a substance as it crystallises, for example in copper(II) sulfate pentahydrate ($CuSO_4 \cdot 5H_2O$).

8.1 Acids and alkalis

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Acids

Acids are substances that produce H^+ ions when they are dissolved in water.

Acids are defined as proton (H^+) donors.

Chapter 8 Acids Bases And Ph Quia

Richard Bailey



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