



# COVALENT BONDING

## Practice Problems

Do your best to solve each of the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{CaF}_2$
  - $\text{C}_2\text{H}_6$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{MgO}$
  - $\text{C}_2\text{H}_5\text{OH}$
- Write the formula for the following compounds:
  - Carbon tetrachloride
  - Carbon disulfide
  - Carbon monoxide
  - Diphosphorus pentoxide
  - Diphosphorus trioxide
  - Diphosphorus pentasulfide
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### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide.  $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane.  $\text{C}$  with four single bonds to  $\text{H}$ .
- Draw the Lewis structure for acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for ammonia.  $\text{N}$  with three single bonds to  $\text{H}$  and one lone pair.
- Draw the Lewis structure for water.  $\text{O}$  with two single bonds to  $\text{H}$  and two lone pairs.



### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
- Which orbital and sub-orbital are represented by the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$

# Chapter 8 Covalent Bonding Test B Answers

**JL Elias**



**Chapter 8 Covalent Bonding Test B Answers:**

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