

Percent composition, Empirical and Molecular Formula Quiz

Find the percent composition of all of the elements in the following compounds (1 pt per line)



$$\begin{array}{l} \text{Al } 26.982 \times 2 = 53.964 \div 294.15 \\ \text{S } 32.065 \times 3 = 96.195 \div 294.15 \\ \text{O } 15.999 \times 9 = 143.991 \div 294.15 \\ \hline 294.15 \end{array}$$

$$\begin{array}{l} \text{Al: } \underline{18.35\%} \\ \text{S: } \underline{32.70\%} \\ \text{O: } \underline{48.95\%} \end{array}$$



$$\begin{array}{l} \text{Na } 22.99 \times 1 = 22.990 \div 102.894 \\ \text{Br } 79.904 \times 1 = 79.904 \div 102.894 \\ \hline 102.894 \end{array}$$

$$\begin{array}{l} \text{Na: } \underline{22.34\%} \\ \text{Br: } \underline{77.66\%} \end{array}$$



$$\begin{array}{l} \text{Fe } 55.845 \times 1 = 55.845 \div 115.853 \\ \text{C } 12.011 \times 1 = 12.011 \div 115.853 \\ \text{O } 15.999 \times 3 = 47.997 \div 115.853 \\ \hline 115.853 \end{array}$$

$$\begin{array}{l} \text{Fe: } \underline{48.20\%} \\ \text{C: } \underline{10.37\%} \\ \text{O: } \underline{41.43\%} \end{array}$$



$$\begin{array}{l} \text{C } 12.011 \times 1 = 12.011 \\ \text{O } 15.999 \times 2 = 31.998 \\ \hline 44.009 \end{array}$$

$$\begin{array}{l} \text{C: } \underline{27.29\%} \\ \text{O: } \underline{72.71\%} \end{array}$$

Chapter 8 Quiz 2 11 12 Practice Quiz Answer Key

Spencer A. Rathus, Jeffrey S. Nevid



Chapter 8 Quiz 2 11 12 Practice Quiz Answer Key:

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