

2018

Name: Phyllis, J. Smith LAB NUMBER: 200

Suppose that you had the correct apparatus for the following experiment. In what state would the liquid be at room temperature?

Compound	mp, °C	bp, °C	State
<u>CH<sub>2</sub>Br<sub>2</sub></u>	<u>9.9</u>	<u>134</u>	<u>liquid</u>
<u>CH<sub>2</sub>Cl<sub>2</sub></u>	<u>-15.2</u>	<u>60.5</u>	<u>liquid</u>
<u>CH<sub>2</sub>I<sub>2</sub></u>	<u>120.7</u>	<u>172.5</u>	<u>solid</u>
<u>CH<sub>3</sub>Br</u>	<u>-93.8</u>	<u>58.8</u>	<u>liquid</u>

1. If you had the same conditions as in 1.1, but a mixture of 100 g of CH<sub>2</sub>Br<sub>2</sub> and 100 g of CH<sub>2</sub>Cl<sub>2</sub> would be available, composition of water to yield 100 g of CH<sub>2</sub>Br<sub>2</sub> would be:

$$\frac{100 \text{ g CH}_2\text{Br}_2}{197 \text{ g}} \left| \frac{1 \text{ mole}}{154 \text{ g}} \right| \frac{1 \text{ mole}}{1 \text{ mole}} \left| \frac{100 \text{ g H}_2\text{O}}{18 \text{ g}} \right| = 27.17 \text{ mol H}_2\text{O}$$

2. If you were in the market and wanted the same composition of 100 g of CH<sub>2</sub>Br<sub>2</sub> in the mixture of the solution in 1.1, you would:

$$\text{a) } \frac{100 \text{ g CH}_2\text{Br}_2}{197 \text{ g}} \left| \frac{1 \text{ mole}}{154 \text{ g}} \right| \frac{1 \text{ mole}}{1 \text{ mole}} \left| \frac{100 \text{ g H}_2\text{O}}{18 \text{ g}} \right| = 27.17 \text{ mol H}_2\text{O}$$

$$\text{b) } \frac{100 \text{ g CH}_2\text{Br}_2}{197 \text{ g}} \left| \frac{1 \text{ mole}}{154 \text{ g}} \right| \frac{1 \text{ mole}}{1 \text{ mole}} \left| \frac{100 \text{ g H}_2\text{O}}{18 \text{ g}} \right| = 27.17 \text{ mol H}_2\text{O}$$

27.17 mol  
moles

27.17 mol  
moles

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**R Barnett**



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