

**WORKSHEET**  
**CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1**  
 CHEMISTRY OF METALS AND ALLOYS

**ANSWERS KEY**

**QUESTION 1. CLASSIFY THE ELEMENTS OF PERIODIC TABLE INTO METALS AND NON-METALS.**

1. Classification of metals is established based on properties such as electrical and thermal conductivity, ductility, malleability, etc. Metals are those elements which have positive values of electrical conductivity.
  - (a) write the balanced chemical equation for the reaction of sodium with water.
 
$$2Na + 2H_2O \rightarrow 2NaOH + H_2$$
  - (b) explain the following observations:
    - (i) Sodium reacts with oxygen to form sodium peroxide.
 
$$2Na + O_2 \rightarrow Na_2O_2$$
    - (ii) Sodium reacts with chlorine to form sodium chloride.
 
$$2Na + Cl_2 \rightarrow 2NaCl$$
2. Some typical reactions between s-block elements are written in the table below. Give necessary conditions. Also indicate the nature of reaction (oxidation-reduction and acid-base).
  - (a) write the balanced chemical equation for the reaction of calcium with water.
 
$$Ca + 2H_2O \rightarrow Ca(OH)_2 + H_2$$
  - (b) indicate the nature of the reaction between calcium and water.
 

Calcium reacts with water to form calcium hydroxide and hydrogen gas. This is an oxidation-reduction reaction.
  - (c) indicate the nature of the reaction between calcium and oxygen.
 

Calcium reacts with oxygen to form calcium oxide. This is an oxidation-reduction reaction.
3. Indicate the type of reaction between calcium and fluorine. Write the balanced chemical equation for the reaction.
  - (a) write the balanced chemical equation for the reaction of calcium with fluorine.
 
$$Ca + F_2 \rightarrow CaF_2$$
  - (b) Give the nature of the reaction between calcium and fluorine.
 

Calcium reacts with fluorine to form calcium fluoride. This is an oxidation-reduction reaction.
  - (c) Give the nature of the reaction between calcium and fluorine.
 

Calcium reacts with fluorine to form calcium fluoride. This is an oxidation-reduction reaction.

**QUESTION 2. CLASSIFY THE ELEMENTS INTO METALS AND NON-METALS.**

1. The metal elements are those elements which have positive values of electrical conductivity.
  - (a) write the balanced chemical equation for the reaction of sodium with water.
 
$$2Na + 2H_2O \rightarrow 2NaOH + H_2$$
  - (b) explain the following observations:
    - (i) Sodium reacts with oxygen to form sodium peroxide.
 
$$2Na + O_2 \rightarrow Na_2O_2$$
    - (ii) Sodium reacts with chlorine to form sodium chloride.
 
$$2Na + Cl_2 \rightarrow 2NaCl$$

# Chem 21 Lab Answers

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