

## Chapter Test II

### A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

#### Column I

- \_\_\_\_\_ 1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
- \_\_\_\_\_ 2. The number of protons in the nucleus of an atom of an element.
- \_\_\_\_\_ 3. The number of electrons in a neutral atom.
- \_\_\_\_\_ 4. The ratio of the actual weight of an element to the standard weight of 16.
- \_\_\_\_\_ 5. The number of atoms of an element in a molecule of a compound.
- \_\_\_\_\_ 6. An element whose atoms have a mass number of 23 and an atomic number of 11.

#### Column II

- a. atomic weight
- b. limiting reagent
- c. molecular weight
- d. stoichiometry
- e. percent yield
- f. atomic number

### B. Multiple Choice

Choose the correct answer for each question.

- \_\_\_\_\_ 1. How many molecules of  $\text{H}_2\text{O}$  are produced when 100 g of  $\text{H}_2$  reacts with 100 g of  $\text{O}_2$ ?
 

<ol style="list-style-type: none"> <li>a. 1000 molecules</li> <li>b. 10000 molecules</li> <li>c. 100000 molecules</li> <li>d. 1000000 molecules</li> </ol>	<ol style="list-style-type: none"> <li>e. 10000000 molecules</li> <li>f. 100000000 molecules</li> <li>g. 1000000000 molecules</li> <li>h. 10000000000 molecules</li> </ol>
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- \_\_\_\_\_ 2. Which of the following is not a chemical change?
  - a.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
  - b.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
  - c.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
  - d.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2$
- \_\_\_\_\_ 3. A 100-gram sample of  $\text{H}_2\text{O}$  is decomposed into hydrogen and oxygen gases. How many grams of hydrogen gas are produced?
 

<ol style="list-style-type: none"> <li>a. 100 g</li> <li>b. 11.1 g</li> <li>c. 111 g</li> <li>d. 11.11 g</li> </ol>	<ol style="list-style-type: none"> <li>e. 111.1 g</li> <li>f. 11.11 g</li> <li>g. 11.111 g</li> <li>h. 11.1111 g</li> </ol>
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# Chemistry Chapter 12 Stoichiometry Test

**Lauren Gardner**



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