

1. A 33.7-g sample of HF is dissolved in water to give 2.0×10^2 mL of solution. The concentration of the solution is:

- A) 1.68 M
B) 0.34 M
C) 0.17 M
D) 8.4 M
E) 16.9 M

$$[HF] = \frac{33.7g HF}{2.0 \times 10^2 mL} \times \frac{1000 mL}{1 L} \times \frac{1 mol HF}{20.01g HF} = 8.4 M$$

2. What mass of calcium chloride, $CaCl_2$, is needed to prepare 3.419 L of a 1.56 M solution?

- A) 243 g
B) 5.33 g
C) 50.6 g
D) 111 g
E) 592 g

$$\frac{3.419 L}{1} \times \frac{1.56 mol CaCl_2}{L} \times \frac{110.98 g CaCl_2}{1 mol CaCl_2} = 592 g CaCl_2$$

3. A 51.24-g sample of $Ba(OH)_2$ is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar $Ba(OH)_2$?

- A) 403 mL**
B) 248 mL
C) 24.8 mL
D) 83.0 mL
E) 331 mL

$$[Ba(OH)_2] = \frac{51.24g Ba(OH)_2}{1.205 L} \times \frac{1 mol Ba(OH)_2}{171.36g Ba(OH)_2} = 0.2481 M$$

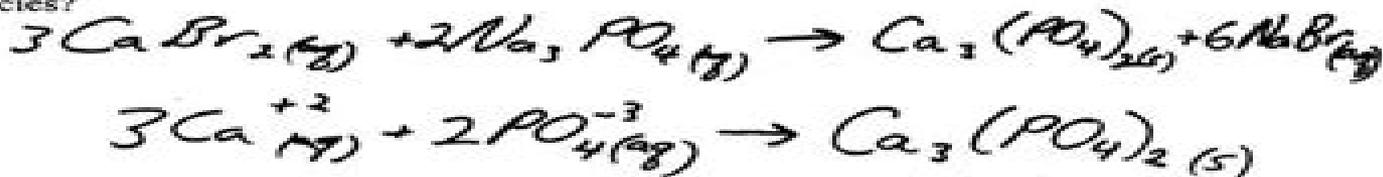
$$M_1 = 0.2481 M \quad M_2 = 0.1000 M$$

$$V_1 = 1.205 L \quad V_2 = ?$$

$$V_2 = \frac{0.1000 M \times 1.000 L}{0.2481 M} = 0.4031 L = 403 mL$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?

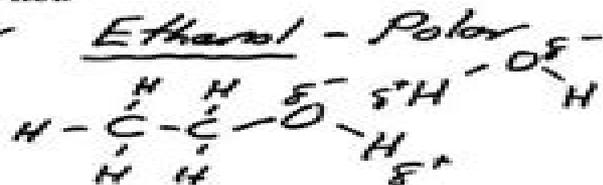
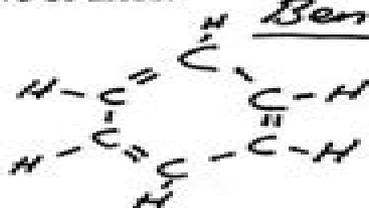
- A) $Ca^{2+}(aq)$
B) $PO_4^{3-}(aq)$
C) $2Ca_3(PO_4)_2(s)$
D) $6NaBr(aq)$
E) $3Ca^{2+}(aq)$



5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?

- A) They have different molar masses.
B) One is ionic, the other is not.
C) One is an electrolyte, the other is not.
D) Ethanol contains a polar O—H bond, and benzene does not.
E) Two of these.

* Know the general solubility rule of "Likes Dissolve Likes"



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Edward Wallace Johnson



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