

# CHEMISTRY

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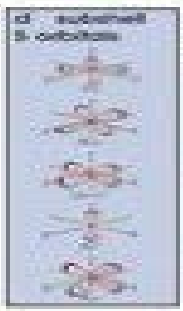
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## PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups and subshells. The groups are numbered 1 through 10. The subshells are labeled as follows:

- 1s subshell:** Group 1 (Hydrogen)
- 2s subshell:** Groups 2 and 3 (Lithium, Beryllium, Boron, Carbon, Nitrogen, Oxygen, Fluorine, Neon)
- 3s subshell:** Groups 4, 5, 6, and 7 (Sodium, Magnesium, Aluminum, Silicon, Phosphorus, Sulfur, Chlorine, Argon)
- 4s subshell:** Groups 8, 9, and 10 (Potassium, Calcium, Scandium, Titanium, Vanadium, Chromium, Manganese, Iron, Cobalt, Nickel, Copper, Zinc, Gallium, Germanium, Arsenic, Selenium, Bromine, Krypton)



## ATOMIC STRUCTURE

Atomic Number,  $Z$ , of an element is the number of protons in the nucleus. For a neutral atom,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of protons and neutrons in the nucleus. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

## ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number,  $n$ , is the number of shells.  $n = 1, 2, 3, \dots$ . The number of orbitals in a shell is  $n^2$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ .

## NUCLEAR CHEMISTRY

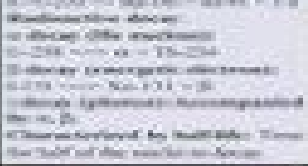
Nuclear reactions occur when particles (neutrons, protons, alpha particles, beta particles, gamma rays) are emitted or absorbed. The mass number,  $A$ , is the sum of protons and neutrons. The atomic number,  $Z$ , is the number of protons. The mass number,  $A$ , is the sum of protons and neutrons. The atomic number,  $Z$ , is the number of protons.

## MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in many-electron atoms. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy.

n	l	m <sub>l</sub>	m <sub>s</sub>	Orbital
1	0	0	±1/2	1s
2	0	0	±1/2	2s
2	1	0	±1/2	2p <sub>z</sub>
2	1	±1	±1/2	2p <sub>x</sub> , 2p <sub>y</sub>
3	0	0	±1/2	3s
3	1	0	±1/2	3p <sub>z</sub>
3	1	±1	±1/2	3p <sub>x</sub> , 3p <sub>y</sub>
3	2	0	±1/2	3d <sub>z<sup>2</sup></sub>
3	2	±1	±1/2	3d <sub>xy</sub> , 3d <sub>xz</sub> , 3d <sub>yz</sub>

Atomic configuration: The arrangement of electrons in an atom. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy.



# Chemistry Quickstudy Reference Guides Academic

**Siddappa N.Byrareddy**



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laminated guide is intended to serve as a reminder and quick study review for all aspects of the chemistry lab from safety to chemicals to instruments and safety again first last and always Designed to find answers fast and to be spill proof it s perfect for high school and early college students studying chemistry as well as anyone interested in learning how scientists safely study the molecules that make up everything around us 6 page laminated guide includes Lab Safety Training General Lab Guidelines Working With Chemicals Exposure to Chemicals First Aid Know Your Lab Reagents Chemical Spills Waste Management Useful Chemical Information Periodic Table Included Lab Must Knows Safe Use of Lab Equipment Data Manipulation Preparing a Solution Inorganic Chemistry Mark Jackson,2010-05-31 Our 3 panel 6 page guide is jam packed with chemistry information not covered in organic chemistry and biochemistry Perfect for science students who want to focus on the aspects of chemistry that go beyond what is covered in material that deals with organic topics this guide covers the key concepts principles figures and formulas that inorganic chemistry students will need to know in order to succeed Tables images and graphic elements further enhance the text Chemistry 2 BarCharts, Inc,Jane P. Gardner,2016 **QuickStudy Academic** Mark Jackson (PhD and Science Professor),2015 Quick Reference for the core essentials of a subject and class that is challenging at best and that many students struggle with In 6 laminated pages our experienced chemistry author and professor gathered key elements organized and designed to use along with your text and lectures as a review before testing or as a memory companion that keeps key answers always at your fingertips AP Chemistry BarCharts, Inc.,2010-12-31 Preparation for the AP Chemistry exam has never been easier once BarCharts gets involved All the latest information regarding this important exam is featured in our jam packed 3 panel guide including sample problems that tackle each area of the test Tables graphs and full color charts and diagrams are also featured to help give students the winning edge they need Periodic Table Basic Mark Jackson,2018 Physics Terminology BarCharts, Inc,Mark Jackson,2013-12-31 Essential terminology for Physics study pulled together in one guide to maximize success in College and High School courses Succinct definitions by our resident Harvard Ph D Chemistry author and professor ensure the usefulness of this handy guide from high school to college Biology ,1999 Algebra S. B. Kizlik,2006-11-30 Essence of the course problems answers formulas equations full color graphics Cover **Periodic Table Basic** Mark Jackson,2002-07-16 Periodic table of the elements Physical Science - Physics and Chemistry Jane Gardner,2022-11 Essential quick reference for a physical science class covering physics and chemistry which would generally be high school level or if combined with our Physical Science 2 guide together would cover a college level Physical Science course Expertly and succinctly written by author STEM curriculum developer and professor Jane Parks Gardner MSc MScEd in an easily accessible outline format with color coded sections This 6 page laminated guide is overflowing with the facts you need to know to boost confidence test scores and grades The value at this price is unmatched for a study tool that can be a review and refresher through both high school and college 6 page laminated guide includes Scientific Method Measurement Significant Figures Precision vs Accuracy Density Matter Physical

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