

Name Morgan WhiteSection 1B1

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1. 1.75 moles of H_2O_2 were placed in a 2.50 L reaction chamber at 307°C . After equilibrium was reached, 1.20 moles of H_2O_2 remained. Calculate the equilibrium constant, K_c , for the reaction.



2. Phosgene, COCl_2 , a poisonous gas, decomposes according to the equation



Calculate K_p for this reaction if $K_c = 0.083$ at 900°C .

$$K_p = K_c(\text{PT})^{\Delta n}$$

3. Calculate K_c for the reaction



given that the concentrations of each species at equilibrium are as follows: $[\text{HI}] = 0.85 \text{ mol/L}$, $[\text{I}_2] = 0.60 \text{ mol/L}$, $[\text{H}_2] = 0.27 \text{ mol/L}$.

4. Consider this reaction at equilibrium:



If the volume of the system is expanded at constant temperature, what change will occur in the position of the equilibrium?

Chemistry 12 Equilibrium Lab Report Answers

Jessica J Manson



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