

## Chapter Test II

### A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

#### Column I

- \_\_\_\_\_ 1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
- \_\_\_\_\_ 2. The number of protons in an atom of an element. It is equal to the number of electrons.
- \_\_\_\_\_ 3. The number of electrons in an atom of an element.
- \_\_\_\_\_ 4. The ratio of the actual weight of an element to the standard weight of 12 grams.
- \_\_\_\_\_ 5. The number of atoms of an element in a molecule of a compound.
- \_\_\_\_\_ 6. An element whose atoms have a mass of 12 units.

#### Column II

- a. atomic weight
- b. limiting reagent
- c. molecular weight
- d. stoichiometry
- e. atomic mass
- f. atomic number
- g. carbon isotope

### B. Multiple Choice

Choose the correct answer for each question.

- \_\_\_\_\_ 1. How many molecules of  $\text{CO}_2$  are produced when 100 g of  $\text{CaCO}_3$  is heated?
 

<ol style="list-style-type: none"> <li>a. <math>10^{23}</math> molecules</li> <li>b. 1000 molecules</li> <li>c. 10000 molecules</li> <li>d. 100000 molecules</li> </ol>	<ol style="list-style-type: none"> <li>e. 1000000 molecules</li> <li>f. 10000000 molecules</li> <li>g. 100000000 molecules</li> <li>h. 1000000000 molecules</li> </ol>
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- \_\_\_\_\_ 2. Which of the following is not a chemical change?
  - a.  $\text{H}_2$  reacts with  $\text{O}_2$  to form  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ .
  - b.  $\text{H}_2$  reacts with  $\text{O}_2$  to form  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ .
  - c.  $\text{H}_2$  reacts with  $\text{O}_2$  to form  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ .
  - d.  $\text{H}_2$  reacts with  $\text{O}_2$ .
- \_\_\_\_\_ 3. Which of the following is not a chemical change?
  - a.  $\text{H}_2$  reacts with  $\text{O}_2$  to form  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ .
  - b.  $\text{H}_2$  reacts with  $\text{O}_2$  to form  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ .
  - c.  $\text{H}_2$  reacts with  $\text{O}_2$  to form  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ .
  - d.  $\text{H}_2$  reacts with  $\text{O}_2$ .

# Chemistry Chapter 12 Test B Stoichiometry

**John Moore**



## **Chemistry Chapter 12 Test B Stoichiometry:**

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**Holt Chemistry** Ralph Thomas Myers, 2004

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