

QUALITATIVE ANALYSIS

Experiment	Observation	Inference
<u>Preliminary Test</u>		
1. Noted the colour of the given salt	<ol style="list-style-type: none"> 1. Blue or Bluish Green 2. Greenish 3. Light Green 4. Dark Brown 5. Light Pink 6. Colourless 	<p>Cu^{2+} or Ni^{2+} Ni^{2+} Fe^{2+} Fe^{3+} Mn^{2+} Absence of Cu^{+2}, Ni^{2+}, Fe^{2+}, Fe^{3+}, Mn^{2+}</p>
2. Noted the solubility	<ol style="list-style-type: none"> 1. Soluble in water 2. Soluble in dilute HCl 	<p>NH_4^+ may be present Absence of group I cations</p>
3. Dry Heating Test: Heated a pinch of the salt in a dry Test tube	<ol style="list-style-type: none"> 1. Colour, odourless gas 2. Colourless gas with rotten egg smell 3. Colourless gas with vinegar smell 4. Colourless gas with ammonia smell 5. Reddish brown gas 6. Greenish yellow gas 7. Violet gas 	<p>CO_3^{2-} may be present S^{2-} may be present CH_3COO^- may be present NH_4^+ may be present NO_3^- or Br^- may be present Cl^- may be present I^- may be present</p>
<u>Identification of acid radical (Anion)</u>		
1. Dilute Acid Test: Treated a pinch of the substance With dilute H_2SO_4	<ol style="list-style-type: none"> 1. Colourless, odourless gas with brisk effervescence which turned lime water milky. 2. Vinegar Smell 3. No characteristic Reaction 	<p>CO_3^{2-} present $\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$</p> <p>$\text{CH}_3\text{COO}^-$ Present $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$</p> <p>Absence of CO_3^{2-}, CH_3COO^-</p>
2. Concentrated Acid Test: Treated a pinch of the substance With concentrated H_2SO_4	<ol style="list-style-type: none"> 1. Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in NH_4OH was brought near the mouth of the test tube 2. Reddish brown vapours with Pungent smell. Solution turns red 3. Deep violet vapours with pungent smell. Sides of test tube 	<p>Cl^- present $\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$ $\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$</p> <p>$\text{Br}^-$ present $\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$ $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$</p> <p>$\text{I}^-$ present $\text{I}^- + \text{H}^+ \rightarrow \text{HI}$ $2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$</p>

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Marcel A. Müller



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