

Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility:** the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution:** a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible:** describes liquids that dissolve in each other. **immiscible:** describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No, if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47. 5.55×10^2 g AgNO_3
48. Solubility increases with pressure.
49. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
50. Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.5M KCl
b. 5.3×10^{-2} M MgCl_2
52. 2.00×10^2 mol
53. a. 5.0×10^{-2} mol NaCl, 20 g NaCl
b. 1.0 mol KNO_3 , 1.0×10^2 g KNO_3
c. 2.5×10^2 mol CaCl_2 , 2.0 g CaCl_2
54. a. 2.3×10^2 g NaCl
b. 2.0 g MgCl_2
55. a. 10% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO_3
c. 0.100M MgCl_2
58. The effective molality of the $\text{Ca}(\text{NO}_3)_2$ solution is 3m. The effective molality of the NaCl solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H_2O to 32.0 g $\text{C}_2\text{H}_5\text{OH}$.
63. a. 100.26°C
b. 101.54°C
64. a. -4.40°C
b. -2.2°C
65. a. -1.1°C
b. -0.24°C
c. -1.3°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

Prentice Hall Chemistry Chapter 13 Assessment Answers

Larry W. Mays



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Chemical Property Estimation Edward Baum, 2018-05-11 Our world is widely contaminated with damaging chemicals and companies create thousands of new potentially dangerous chemicals each year Due to the difficulty and expense of obtaining accurate measurements and the unreliability of reported values we know surprisingly little about the properties of these contaminants Determining the properties of chemicals is critical to judging their impact on environmental quality and in making decisions about emission rates clean up and other important public health issues *Chemical Property Estimation* describes modern methods of estimating chemical properties methods which cost much less than traditional laboratory techniques and are sufficiently accurate for most environmental applications Estimation methods are used to screen chemicals for testing design monitoring and analysis methods design clean up procedures and verify experimental measurements The book discusses key methods for estimating chemical properties and considers their relative strengths and weaknesses Several chapters are devoted to the partitioning of chemicals between air water soil and biota and properties such as solubility vapor pressure and chemical transport Each chapter begins with a review of relevant theory and background information explaining the applications and limitations of each method Sample calculations and practical advice on how and when to use each method are included as well Each method is evaluated for accuracy and reliability Computer software databases and internet resources are evaluated as well as other supplementary material such as fundamental constants units of measure and more **Introductory Chemistry** Nivaldo J. Tro, 2006 Sm Fundamentals Chemistry T/I Burns, 1999-03 *High Temperature Corrosion* César A. C. Sequeira, 2018-12-14 Reviews the science and engineering of high temperature corrosion and provides guidelines for selecting the best materials for an array of system processes High temperature corrosion HTC is a widespread problem in an array of industries including power generation aerospace automotive and mineral and chemical processing to name a few This book provides engineers physicists and chemists with a

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