

## Topic: COLLIGATIVE PROPERTIES OF NONELECTROLYTES AND ELECTROLYTE SOLUTIONS

**Colligative properties** – are properties of a solution that depend only on the number and not on the identity of the solute particles.

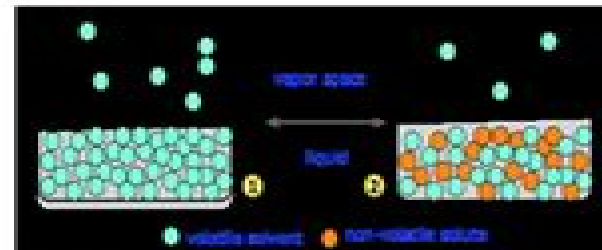
- vapor pressure lowering
- boiling point elevation and
- freezing point depression

### Effect of solute concentration on the colligative properties of solutions

The concentration or amount of nonvolatile solute (i.e., a solute that does not have a vapor pressure of its own) in the solution has an effect on the colligative properties of solutions. The effect would depend on the ratio of the number of particles of solute and solvent in the solution and not on the identity of the solute. However, it is necessary to take into account whether the solute is an electrolyte or a nonelectrolyte.

### Effects of electrolyte and nonelectrolyte on colligative properties solutions.

**1. Vapor Pressure Lowering** – When a nonvolatile solute is added to a pure solvent, the vapor pressure of the solution is lower than that of the pure solvent. This is because the solute particles interfere with the solvent molecules' ability to escape into the vapor phase. The more solute particles present, the greater the interference and the lower the vapor pressure. This is a colligative property because it depends only on the number of solute particles, not their identity.



**2. Boiling point Elevation** – The addition of a nonvolatile solute lowers the vapor pressure of the solution; consequently the temperature must be raised to restore the vapor pressure of the solution to the value conforming to the pure solvent. Specifically, the temperature at which the vapor pressure is 1 atm will be higher than the normal boiling point by an amount known as the boiling point elevation.

**The lowering of the vapor pressure in a solution causes the boiling point of the solution to be higher than pure solvent**



**Boiling point elevation is a colligative property that depends on the number of solute particles in the solution, not their identity.**

# Properties Of Solutions Electrolytes And Nonelectrolytes Lab Report

**CO Houle**



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