

Chapter 6 Thermochemistry

Homework Problems

- The enthalpy change for the fermentation (without oxygen, O_2 gas) reaction of glucose to alcohol and carbon dioxide is -67 kJ.
 - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
 - Is the reaction exothermic or endothermic?
 - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is 0.71 J/ $^{\circ}C$ g. Calculate the energy required to:
 - Raise the temperature of 1.0 mol graphite by $1.0^{\circ}C$.
 - Raise the temperature of 850 g graphite by $150^{\circ}C$.
 - Raise the temperature of 75 kg graphite from 294 K to 348 K.
- In a coffee-cup calorimeter, 100.0 mL of 1.0 M NaOH and 100.0 mL of 1.0 M HCl are mixed. Both solutions were originally at $24.6^{\circ}C$. After the reaction, the final temperature is $31.3^{\circ}C$. Assume that all the solutions have a density of 1.0 g/mL and a specific heat capacity of 4.18 J/g $^{\circ}C$. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ($C_{15}H_{26}O$) has an energy of combustion of -5903.6 kJ/mol. A sample of camphor with a mass of 0.1204 g is burned in a bomb calorimeter, the temperature increases by $2.28^{\circ}C$. Calculate the heat capacity of the calorimeter.
A 0.1964 g sample of quinone ($C_6H_4O_2$) is burned in a bomb calorimeter that has a heat capacity of 1.56 kJ/ $^{\circ}C$. The temperature of the calorimeter increases by $3.2^{\circ}C$. Calculate the energy of combustion of quinone per gram and per mole.

Chapter 6 Thermochemistry Homework Problems

Lingsheng Yao



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