Chapter 8: Covalent Bonding

Students should:

1.14 know what is meant by the term molecule

1.44 know that a covalent bond is formed between atoms by the sharing of a pair of electrons

1.45 understand covalent bonds in terms of electrostatic attractions

1.46 understand how to use dot-and-cross diagrams to represent covalent bonds in:

- diatomic molecules, including hydrogen, oxygen, nitrogen, halogens and hydrogen halides
- inorganic molecules including water, ammonia and carbon dioxide
- organic molecules containing up to two carbon atoms, including methane, ethane, ethene and those containing halogen atoms

1.47 explain why substances with simple molecular structures are gases or liquids, or solids with low melting and boiling points

the term intermolecular forces of attraction can be used to represent all forces between molecules

1.48 explain why the melting and boiling points of substances with simple molecular structures increase, in general, with increasing relative molecular mass

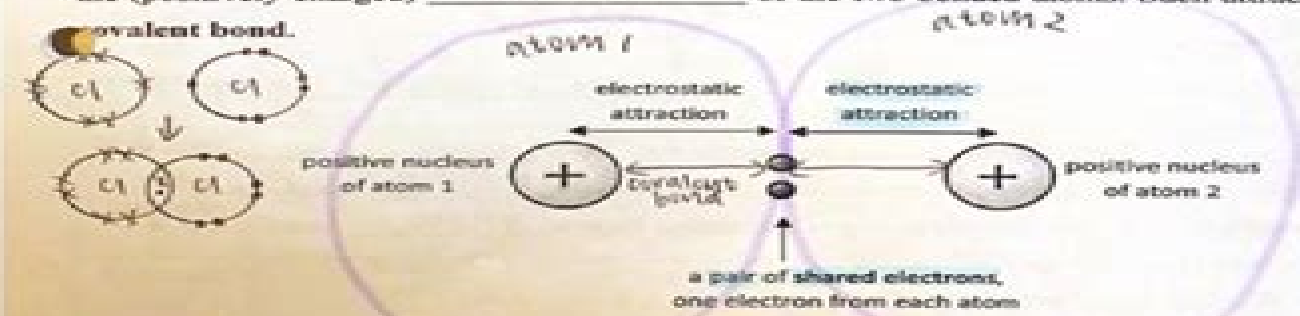
1.49 explain why substances with giant covalent structures are solids with high melting and boiling points

1.50 explain how the structures of diamond, graphite and C_{60} fullerene influence their physical properties, including electrical conductivity and hardness

1.51 know that covalent compounds do not usually conduct electricity

8.1 Covalent bonding

A covalent bond is resulted when non-metal atoms combine with non-metal atoms by sharing of outermost shell electrons. Strong, directional electrostatic attraction is formed between the (negatively charged) shared electrons and the (positively charged) nuclei of the two bonded atoms. Such attraction is called a covalent bond.



The bonded atoms can then attain an octet or duplet of electrons and become stable (i.e. each atom attains the stable electronic structure of its nearest noble gas).

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