

Assessment

Chapter 8 Practice Test**Chapter: Chemical Equations and Reactions**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. You mix solution A with solution B in a beaker. Which of the following observations does *not* help you prove that a chemical reaction has occurred?
- The beaker becomes warm.
 - You see gas bubble out of solution.
 - Solution A, solution B, and the mixture are all yellow.
 - None of the above; all help prove that a chemical reaction has occurred.
- _____ 2. In the reaction described by the word equation
sodium + water \rightarrow sodium hydroxide + hydrogen gas
- sodium hydroxide is a product.
 - sodium is a product.
 - water is a reactant.
 - Both (a) and (c)
- _____ 3. Which of the following indicates that a chemical equation is balanced?
- The numbers of atoms of each element are the same on both sides of the equation.
 - All of the coefficients are the same.
 - The numbers of molecules on each side of the equation are equal.
 - The sums of the coefficients on each side of the equation are equal.
- _____ 4. From a complete and correctly written chemical equation, you can obtain the
- chemical formulas of the reactants and products.
 - relative amounts of the reactants and products.
 - physical states of the reactants and products.
 - All of the above
- _____ 5. To balance a chemical equation, you must
- adjust the subscripts.
 - adjust the number of products.
 - adjust the number of reactants.
 - adjust the coefficients.

Chapter 8 Test 16

Lawrence Alexander Sharp



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