1. A 33.7-g sample of HF is dissolved in water to give 2.0×10^2 mL of solution. The concentration of the solution is:
   A) 1.68 M
   B) 0.34 M
   C) 0.17 M
   D) 8.4 M
   E) 16.9 M

$$[HF] = \frac{33.7\text{g HF}}{2.0\times10^2\text{mL}} \times \frac{1000\text{mL}}{1\text{L}} \times \frac{1\text{mol HF}}{20.01\text{g HF}} =$$

$$[HF] = 8.4\text{ M}$$

2. What mass of calcium chloride, $CaCl_2$, is needed to prepare 3.419 L of a 1.56 M solution?
   A) 243 g
   B) 5.33 g
   C) 50.6 g
   D) 111 g
   E) 592 g

$$\overbrace{\frac{3.419\text{L}}{1} \times \frac{1.56\text{mol CaCl}_2}{\text{L}}}^{\text{mole CaCl}_2} \times \frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} = 592\text{g CaCl}_2$$

3. A 51.24-g sample of $Ba(OH)_2$ is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar $Ba(OH)_2$?
   A) 403 mL
   B) 248 mL
   C) 24.8 mL
   D) 83.0 mL
   E) 331 mL

$$[Ba(OH)_2] = \frac{51.24\text{g Ba(OH)}_2}{1.205\text{L}} \times \frac{1\text{mol Ba(OH)}_2}{171.36\text{g Ba(OH)}_2}$$

$$[Ba(OH)_2] = 0.2481\text{ M}$$

$M_1 = 0.1000\text{M}$ $M_2 = 0.2481\text{M}$

$V_1 = 1.000\text{L}$ $V_2 = ?$

$$V_2 = \frac{0.1000\text{M} \times 1.000\text{L}}{0.2481\text{M}}$$

$$V_2 = 0.4031\text{L} = 403\text{mL}$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?
   A) $Ca^{2+}(aq)$
   B) $PO_4^{3-}(aq)$
   C) $2Ca_3(PO_4)2(s)$
   D) 6NaBr(aq)
   E) $3Ca^{2+}(aq)$

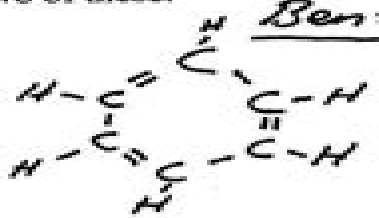
$$3CaBr_{2(aq)} + 2Na_3PO_{4(aq)} \rightarrow Ca_3(PO_4)_{2(s)} + 6NaBr_{(aq)}$$

$$3Ca^{+2}_{(aq)} + 2PO_{4(aq)}^{-3} \rightarrow Ca_3(PO_4)_{2(s)}$$

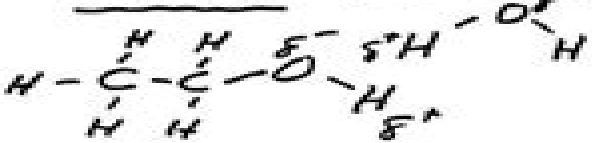
5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?
   A) They have different molar masses.
   B) One is ionic, the other is not.
   C) One is an electrolyte, the other is not.
   D) Ethanol contains a polar O—H bond, and benzene does not.
   E) Two of these.

* Know the general solubility rule of "Likes Dissolve Likes"

Benzene – Non Polar



Ethanol – Polar



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