

ANSWERS TO CONVERSION FACTOR PROBLEMS

Conversion factors can be used to convert units or to convert between equivalent ways of expressing a quantity. The quantity in the problem is multiplied by one or more "conversion factors," in which the numerator is equal to the denominator. Since the numerator and denominator of the conversion factor are equal, multiplying by the conversion factor is like multiplying by 1 and thus does not change the value of the original quantity. Use the table of English to Metric equivalents as needed. All answers should be in *significant figures*!

A) Problems with a single conversion factor.

1. How many cm are in 18.9 inches?

$$18.9 \text{ in} \times \frac{2.54 \text{ cm}}{1 \text{ in}} = 47.5 \text{ cm}$$

2. How many grams are in 0.143 ounces?

$$0.143 \text{ oz} \times \frac{1 \text{ gram}}{0.03527 \text{ oz}} = 4.05 \text{ grams}$$

B) Problems with two or more conversion factors.

1. How many mL are in 0.037 quarts?

$$0.037 \text{ qt} \times \frac{1 \text{ L}}{1.057 \text{ qt}} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 35 \text{ mL}$$

2. How many grams are in 0.397 pounds (lbs)?

$$0.397 \text{ lb} \times \frac{1 \text{ kg}}{2.20 \text{ lb}} \times \frac{1000 \text{ g}}{1 \text{ kg}} = 180 \text{ g} = 1.80 \times 10^3 \text{ g}$$

3. How many micrograms are in 6.8×10^{-7} ounces?

$$6.8 \times 10^{-7} \text{ oz} \times \frac{1 \text{ g}}{0.0353 \text{ g}} \times \frac{1000 \text{ mg}}{1 \text{ g}} \times \frac{1000 \text{ } \mu\text{g}}{1 \text{ mg}} = 19 \text{ } \mu\text{g}$$

C) Using density as a conversion factor. The density of Al is 2.70 g/cm^3 .

Remember, $1 \text{ mL} = 1 \text{ cm}^3$.

1. What is the volume of 0.810 grams of Al?

$$0.810 \text{ g} \times \frac{1 \text{ cm}^3}{2.70 \text{ g}} = 0.300 \text{ cm}^3 \text{ (or } 0.300 \text{ mL)}$$

2. Find the mass in grams of 0.327 liters of Al.

$$0.327 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} \times \frac{2.70 \text{ g}}{1 \text{ cm}^3} = 883 \text{ g}$$

Chemistry Conversion Factors Practice Problems With Solutions

Margaret G. Kimble



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connections that exist between these three branches of chemistry and allows instructors to discuss these interrelationships while the material is still fresh in students minds This text is an unbound binder ready edition Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Essentials of Introductory Chemistry* Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry *Introductory Chemistry* Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students **Chemistry** George Tyler Miller, 1987 *Chemistry* John McMurry, Robert C. Fay, 1998 A full year course taken primarily by chemistry majors other science majors especially biology and pre health and engineering students First introduced in 1995 McMurry Fay s Chemistry is now recognized as one of the leading books in science education The second edition refines the qualities that led to the text s success in the first place The text is a beautifully presented and well written general chemistry text The chapter on gas laws where the combination of narrative and illustrations leads the students to almost derive the Kinetic Molecular theory on their own Mildred Hall Clark State Community College **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the

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General Chemistry, Reactions First Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

A Learning Program for Chemistry Peter Demmin, 1993-11 *Improving Student Comprehension in Chemistry Laboratories* Tracy Lynn Haroff, 2006

Chemistry Karen Timberlake, 1983 **Problem Solving in General Chemistry** Ronald A. DeLorenzo, 1981 **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994

Fundamentals of Chemistry Ralph A. Burns, 2003 For one semester preparatory chemistry courses or general purpose introductory chemistry courses This clearly written well illustrated versatile textbook provides thorough coverage of chemistry with a balance of problem solving skills real world applications and an emphasis on critical thinking and the process of science A supporting theme throughout the text continually emphasizes that chemistry is everywhere

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