

QUESTION 1

Volume of pipette used 25.00 cm^3
 (V₀)

Titration	Rough	1	2	3
Final burette reading (cm ³)	20.80	41.40	20.40	40.90
Initial burette reading (cm ³)	0.00	20.80	0.00	20.40
Volume of A used	20.80	20.60	20.40	20.50

② ^{average} Volume of A used = $\frac{20.60 + 20.40 + 20.50}{3}$
 $= 20.50 \text{ cm}^3$

③.1) Concentration of A in mol/dm³ = $\frac{\text{conc. in g/dm}^3}{\text{molar mass}}$
 $= \frac{4.4 \text{ g/dm}^3}{2[13] + 32 + 4[16]}$
 $= \frac{4.4}{98} = 0.0449 \text{ mol/dm}^3$

③.2) Concentration of B in mol/dm³, C_B

Using $\frac{C_A V_A}{C_B V_B} = \frac{n_A}{n_B}$

$$\frac{0.0449 \times 20.50}{C_B \times 25} = \frac{1}{1}$$

$$C_B = \frac{0.0449 \times 20.50}{25} = 0.03287 \text{ mol/dm}^3$$

Chemistry Practical Question And Answer

Patrick Vollmar



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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Chemistry Practical Question And Answer** a literary masterpiece penned by way of a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those that partake in its reading experience.

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