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EXPERIMENT 3 Half-Wave and Full-Wave Rectification

Objective

To calculate, compare, draw, and measure the DC output voltages of half-wave and full-wave rectifier circuits.

Tools and Equipments Required

DMM (Digital Multi Meter)	
DC Power Supply	
Function Generator	
2.2 k Ω	x 3
Silicon Diode	x 4

Theory and Descriptions

The primary function of half-wave and full-wave rectification systems is to establish a DC level from a sinusoidal input signal that has zero average (DC) level.

The half-wave voltage signal of Fig.2.1, normally established by a network with a single diode, has an average or equivalent DC voltage level equal to 31.8% of the peak voltage V_m .

That is
$$V_{dc} = 0.318 \cdot V_{peak} \text{ volts} \quad (\text{half-wave}) \quad \text{eq.2.1}$$

The full-wave rectifier signal of Fig.2.1 has twice the average or DC level of the half-wave signal, or 63.6% of the peak value V_m .

That is
$$V_{dc} = 0.636 \cdot V_{peak} \text{ volts} \quad (\text{full-wave}) \quad \text{eq.2.2}$$

For large sinusoidal inputs ($V_m \gg V_T$) the forward-biased transition voltage V_T of a diode can be ignored. However, for situations when the peak value of the sinusoidal signal is not that much greater than V_T , V_T can have a noticeable effect on V_{DC} .

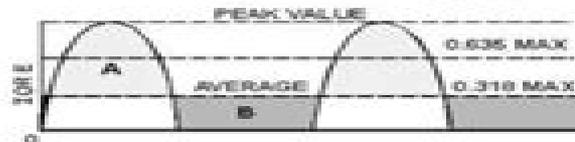


Figure 2.1 half- and full-wave rectified signal

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AW Chikering



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web hess s law labs by austin lee alayna baron lily zmachinski introduction in order to calculate the enthalpy change for the combustion of magnesium oxide $\text{mg} + \frac{1}{2} \text{O}_2 \rightarrow \text{MgO}$ we used

additivity of heats of reaction hess s law lab report - Aug 03 2022

web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows $\text{NaOH} + \text{Na} \rightarrow \text{NaOH} + \text{aq}$

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web hess s laws abstract the following experiment was designed using hess s law and simple calorimetry to determine the enthalpy for NaOH hess s law says that the overall change in enthalpy for the reaction is the sum of all changes regardless of the different stages or phases of a reaction

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web hess s law states that if a reaction is the sum of two or more other reactions the Δh for the overall process must be the sum of the Δh values of the constituent reactions enthalpy change Δh is independent of the path that a reaction follows to move from reactants

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web this activity provides a demonstration of hess law using three reactions the solubility NaOH in water the solubility NaOH in HCl and the reaction of a solution of HCl and a solution of NaOH

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web hess s law worksheet answers 1 calculate h for the reaction $\text{C}_2\text{H}_4 + \text{H}_2 \rightarrow \text{C}_2\text{H}_6$ from the following data $\text{C}_2\text{H}_4 + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O} \quad \Delta h = 1411 \text{ kJ}$ $\text{C}_2\text{H}_6 + \frac{3}{2} \text{O}_2 \rightarrow 2 \text{CO}_2 + 3 \text{H}_2\text{O} \quad \Delta h = 1560 \text{ kJ}$

7a first law enthalpy calorimetry and hess s law worksheet - Oct 05 2022

web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability

to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction c 2 h 4

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web thus for reaction 9 2 in which 1 mole of h₃o and 1 mole of oh combine to form 2 moles of h₂o h 55 8 kj as shown note that the enthalpy change h h_{products} h_{reactants} is positive if heat is absorbed that is if h_{products}

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 naoh s na aq oh aq h 1 2 naoh s h aq cl aq h 2 o na aq cl aq h 2 3