

Stoichiometry Lab: Vinegar and Baking Soda Do Not Write On This Sheet

Purpose: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield.



Materials: Baking Soda (NaHCO_3), Vinegar (CH_3COOH), and 2 plastic cups, scale.

Procedure: Find and record the mass of cup A. With cup A still on the scale, add approximately 10.0 g of baking soda to the cup. **The mass does not have to be exact, as long as you carefully record your results.**

Place cup B on the scale, weigh and record approximately 50.0 g of vinegar. **The mass does not have to be exact, as long as you carefully record your results.**

Slowly add vinegar to cup A until the reaction has stopped. **DO NOT** add all of the vinegar, just enough to complete the reaction. Reweigh and record both cup A and B. Calculate the mass of CO_2 that escaped.

Data:

- a. Mass of Cup A _____ g
- b. Mass of Cup A and baking soda _____ g
- c. Calculate mass of baking soda (b - a) _____ g
- d. Mass of Cup B with vinegar _____ g
- e. Mass of Cup B after reaction _____ g
- f. Calculate mass of vinegar poured into Cup A (d - e) _____ g
- g. Mass of Cup A after reaction _____ g
- h. Calculate mass of product after reaction (g - a) _____ g
- i. Calculate baking soda + vinegar (c + f) _____ g
- j. Calculate mass of CO_2 lost (i - h) _____ g

Baking Soda Stoichiometry Lab Answers

James Joseph Gallagher



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so he lifted his eyes and looked and behold three men were standing by him and when he saw them he ran from the tent door to meet them and bowed himself to the ground 3 and said my lord if i have now

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