

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make use of Le Chatelier's analysis and an equilibrium shift construction through applying Le Chatelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.  
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.  
 $\text{FeSCN}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq}) + \text{H}^{+}(\text{aq})$
4. In part 4, what is the conjugate acid of the base  $\text{FeSCN}^{2+}$ ?  $\text{H}^{+}$  or  $\text{H}_2\text{O}$ . Is  $\text{H}_2\text{O}$  a base or acid?  
Conjugate acid:  $\text{H}^{+}$  is the conjugate acid and  $\text{H}_2\text{O}$  is the conjugate base for  $\text{FeSCN}^{2+}$ . The conjugate acid  $\text{H}^{+}$  has no conjugate base and the conjugate base  $\text{H}_2\text{O}$  has a conjugate acid.
5. What is the conjugate acid of the base  $\text{FeSCN}^{2+}$ ?  $\text{H}^{+}$  or  $\text{H}_2\text{O}$ . Is  $\text{H}_2\text{O}$  a base or acid?  
Conjugate acid:  $\text{H}^{+}$  is the conjugate acid and  $\text{H}_2\text{O}$  is the conjugate base for  $\text{FeSCN}^{2+}$ .
6. What is the conjugate acid of the base  $\text{FeSCN}^{2+}$ ?  $\text{H}^{+}$  or  $\text{H}_2\text{O}$ . Is  $\text{H}_2\text{O}$  a base or acid?  
Conjugate acid:  $\text{H}^{+}$  is the conjugate acid and  $\text{H}_2\text{O}$  is the conjugate base for  $\text{FeSCN}^{2+}$ .
7. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + \text{OH}^{-} \rightleftharpoons \text{FeOH}^{2+}$
8. Write your own name for  $\text{FeOH}^{2+}$ .  
Substituted  $\text{Fe}^{3+}$  with a hydroxide group that produces a substituted metal  $\text{FeOH}^{2+}$  and a hydroxide  $\text{OH}^{-}$  as the anion.
9. What is a substitution reaction?  
A reaction that results in the same form as both the reactants and products substituted molecules. It is reversible and does not affect equilibrium.

# Investigating Chemical Equilibrium Lab 12a Answer

**TD Snyder**



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