

## Solutions

<b>Word</b>	<b>Definition</b>
Aqueous	A solution in which the solvent is water.
Colligative property	A property of a solution that is dependent on concentration. Examples include boiling point, freezing point and vapor pressure.
Molality	The concentration of a solution measured in moles of solute per kilogram of solvent.
Molarity	The concentration of a solution measured in moles of solute per liter of solution.
Parts per million	The concentration of a solution measured in mass of solute per mass of solution multiplied by one million.
Percent by mass	The concentration of a solution measured in mass of solute per mass of solution multiplied by one hundred.
Percent by volume	The concentration of a solution measured in volume of solute per volume of solution multiplied by one hundred.
Saturated	A solution that has the maximum concentration of solute possible in a given quantity of solvent at a given temperature, a solution at equilibrium.
Solubility	The maximum quantity of solute that can be dissolved in a given quantity of solvent at a given temperature to make a saturated solution.
Solute	A substance that is torn apart and kept separate by solvent particles.
Solution	A homogeneous mixture formed from a solute dissolved into a solvent.
Solvent	A substance that attaches to solute particles, tears them apart from each other and keeps them apart.
Supersaturated	A solution in which there is an excess of solute beyond the solubility point for a given temperature. Either the excess will precipitate or it will remain precariously dissolved until the solution is agitated, whereupon the excess will precipitate out.
Unsaturated	A solution in which there are still solvent particles free to attach to and dissolve more solute particles.

# Mark Rosengarten Answers Solutions

**Brendan G. Carr**



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