

Ch. 3: Stoichiometry

Avogadro's Number

- **Mole:** a measuring unit
 - 1 mole of atoms = 6.022×10^{23} atoms

Counting Atoms Using Moles

- A mole is the **exact number of atoms** present in a sample that weighs the element's atomic weight
 - Ex. 1 mole of Fe = 6.022×10^{23} atoms = 55.847 g
 - Ex. What is the mass in grams of 4.30 moles of Al?
Molar mass of Al = 26.98 g/mol

$$4.3 \text{ mol Al} \times \frac{26.98 \text{ g}}{1 \text{ mol Al}} = 116 \text{ g of Al}$$

- The molar mass of a molecule is the **sum of the atomic weight** of all the atoms in a molecule
- **Molecular formula weight:** molar mass of a compound
 - Ex. What is the molar mass of methane gas CH_4 ?
Molar mass of carbon = 12.01 g/mol
Molar mass of hydrogen = 1.01 g/mol

$$\text{Methane} = 12.01 + 4(1.01) = 16.04 \text{ g/mol}$$

Percent Composition by Mass

- **Percent Composition:** the percent by mass, of each element in the compound
- **Formula:**
 - % composition of an element = $\frac{(\# \text{ of atoms of the element}) \times (\text{molar mass of element})}{\text{Molar mass of the compound}} \times 100\%$
 - Ex. Calculate the percent composition by mass of C_3H_8
Molar mass of C_3H_8 = $3(12.01) + 8(1.01) = 44.11 \text{ g/mol}$

$$\% \text{ of C} = \frac{3(12.01)}{44.11} \times 100\% = 81.68\%$$

Empirical and Molecular Formula

- **Empirical Formula:** the simplest whole number ratio of element in a composition
 - Does NOT tell the number of atoms in a molecule

How to find the Empirical Formula from % mass

Part A: Determine Empirical Formula

1. Convert the data from % to grams for each element
2. Convert from grams to moles for each element by dividing atomic mass
3. Divide each value by the smallest of the values and round answers to the nearest whole number

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Table of Contents Mastering Chemistry Answers Chapter 3 Cycamp

1. Understanding the eBook Mastering Chemistry Answers Chapter 3 Cycamp
 - The Rise of Digital Reading Mastering Chemistry Answers Chapter 3 Cycamp
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answers Chapter 3 Cycamp
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answers Chapter 3 Cycamp
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answers Chapter 3 Cycamp
 - Personalized Recommendations
 - Mastering Chemistry Answers Chapter 3 Cycamp User Reviews and Ratings
 - Mastering Chemistry Answers Chapter 3 Cycamp and Bestseller Lists
5. Accessing Mastering Chemistry Answers Chapter 3 Cycamp Free and Paid eBooks
 - Mastering Chemistry Answers Chapter 3 Cycamp Public Domain eBooks
 - Mastering Chemistry Answers Chapter 3 Cycamp eBook Subscription Services

- Mastering Chemistry Answers Chapter 3 Cycamp Budget-Friendly Options
6. Navigating Mastering Chemistry Answers Chapter 3 Cycamp eBook Formats
 - ePub, PDF, MOBI, and More
 - Mastering Chemistry Answers Chapter 3 Cycamp Compatibility with Devices
 - Mastering Chemistry Answers Chapter 3 Cycamp Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answers Chapter 3 Cycamp
 - Highlighting and Note-Taking Mastering Chemistry Answers Chapter 3 Cycamp
 - Interactive Elements Mastering Chemistry Answers Chapter 3 Cycamp
 8. Staying Engaged with Mastering Chemistry Answers Chapter 3 Cycamp
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answers Chapter 3 Cycamp
 9. Balancing eBooks and Physical Books Mastering Chemistry Answers Chapter 3 Cycamp
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answers Chapter 3 Cycamp
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mastering Chemistry Answers Chapter 3 Cycamp
 - Setting Reading Goals Mastering Chemistry Answers Chapter 3 Cycamp
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mastering Chemistry Answers Chapter 3 Cycamp
 - Fact-Checking eBook Content of Mastering Chemistry Answers Chapter 3 Cycamp
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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