

Chemistry Worksheet BUNDLE



Calculating Molar Mass

Instructions:
Calculate the molar mass for each of the compounds below. Make sure to give all units and include your work.



Example: H_2O

H = 1.008 g/mol \times 2 = 2.016 g/mol O = 16.00 g/mol \times 1 = 16.00 g/mol

Compound	Formula	Molar Mass
Hydrogen peroxide	H_2O_2	
Carbon dioxide	CO_2	
Ammonia	NH_3	
Glucose	$\text{C}_6\text{H}_{12}\text{O}_6$	
Acetic acid	$\text{C}_2\text{H}_4\text{O}_2$	
Hydrochloric acid	HCl	
Sulfuric acid	H_2SO_4	
Phosphoric acid	H_3PO_4	
Hydrofluoric acid	HF	
Hydrobromic acid	HBr	
Hydroiodic acid	HI	
Hydrocyanic acid	HCN	
Hydroperchloric acid	HClO_4	
Hydroperbromic acid	HBrO_4	
Hydroperiodic acid	HIO_4	
Hydroperacetic acid	CH_3COOOH	
Hydroperoxyacetic acid	$\text{CH}_3\text{OOCH}_2\text{OOH}$	
Hydroperoxybenzoic acid	$\text{C}_6\text{H}_5\text{COOOH}$	
Hydroperoxyacetic anhydride	$(\text{CH}_3\text{COO})_2\text{O}$	
Hydroperoxybenzoic anhydride	$(\text{C}_6\text{H}_5\text{COO})_2\text{O}$	
Hydroperoxyacetic acid chloride	$\text{CH}_3\text{COOCH}_2\text{OOHCl}$	
Hydroperoxybenzoic acid chloride	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHCl}$	
Hydroperoxyacetic acid bromide	$\text{CH}_3\text{COOCH}_2\text{OOHBr}$	
Hydroperoxybenzoic acid bromide	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHBr}$	
Hydroperoxyacetic acid iodide	$\text{CH}_3\text{COOCH}_2\text{OOHI}$	
Hydroperoxybenzoic acid iodide	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHI}$	
Hydroperoxyacetic acid fluoride	$\text{CH}_3\text{COOCH}_2\text{OOHF}$	
Hydroperoxybenzoic acid fluoride	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHF}$	
Hydroperoxyacetic acid chloride	$\text{CH}_3\text{COOCH}_2\text{OOHCl}$	
Hydroperoxybenzoic acid chloride	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHCl}$	
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Hydroperoxybenzoic acid fluoride	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHF}$	

ANSWER
SHEETS!



Percent Composition

Instructions:
Calculate the percent composition of each element in the compounds below. Show your work.



Example: H_2O

H = 1.008 g/mol \times 2 = 2.016 g/mol O = 16.00 g/mol \times 1 = 16.00 g/mol

Compound	Formula	% H	% O
Hydrogen peroxide	H_2O_2		
Carbon dioxide	CO_2		
Ammonia	NH_3		
Glucose	$\text{C}_6\text{H}_{12}\text{O}_6$		
Acetic acid	$\text{C}_2\text{H}_4\text{O}_2$		
Hydrochloric acid	HCl		
Sulfuric acid	H_2SO_4		
Phosphoric acid	H_3PO_4		
Hydrofluoric acid	HF		
Hydrobromic acid	HBr		
Hydroiodic acid	HI		
Hydrocyanic acid	HCN		
Hydroperchloric acid	HClO_4		
Hydroperbromic acid	HBrO_4		
Hydroperiodic acid	HIO_4		
Hydroperacetic acid	CH_3COOOH		
Hydroperoxyacetic acid	$\text{CH}_3\text{OOCH}_2\text{OOH}$		
Hydroperoxybenzoic acid	$\text{C}_6\text{H}_5\text{COOOH}$		
Hydroperoxyacetic anhydride	$(\text{CH}_3\text{COO})_2\text{O}$		
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Hydroperoxybenzoic acid fluoride	$\text{C}_6\text{H}_5\text{COOCH}_2\text{OOHF}$		

25
PROBLEMS

Calculating Molar Mass and Percent Composition



Molar Mass And Percent Composition Worksheet Answers

RM Cervero



Molar Mass And Percent Composition Worksheet Answers:

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Holt Chemistry Ralph Thomas Myers,2004

Chemistry Homework Frank Schaffer Publications,Joan DiStasio,1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more

Counting Moles Nigel P. Freestone,2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations

Bite-Sized Chemistry Calculations John Lambert,2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those

following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance

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Formulas, Moles and Reactions Alessio Mangoni, 2025-05-28 This comprehensive and rigorous educational text is dedicated to the study and application of the fundamentals of quantitative chemistry Through sixty clearly and precisely solved exercises it addresses essential topics such as molar mass and molecular weight the concept of the mole Avogadro s number the calculation of atoms and molecules in known quantities percent composition of compounds determination of empirical and molecular formulas and the balancing of chemical equations It serves as an effective tool for chemistry students and for anyone wishing to deepen their understanding of the discipline through a practical and methodical approach providing solid skills that are essential for continuing in scientific studies

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