

Chemistry

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(Chapter 11) (Alcohols Phenols and Ethers) (Exercises)

(Class 12)



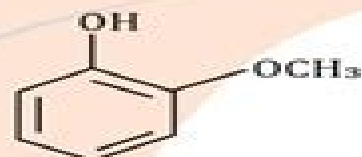
Question 11.15:

Explain why is *ortho* nitrophenol more acidic than *ortho* methoxyphenol?

Answer 11.15:



o - Nitrophenol



o - Methoxyphenol

The nitro-group is an electron-withdrawing group. The presence of this group in the *ortho* position decreases the electron density in the O-H bond. As a result, it is easier to lose a proton. Also, the *o*-nitrophenoxide ion formed after the loss of protons is stabilized by resonance. Hence, *ortho* nitrophenol is a stronger acid.

On the other hand, methoxy group is an electron-releasing group. Thus, it increases the electron density in the O-H bond and hence, the proton cannot be given out easily.

For this reason, *ortho*-nitrophenol is more acidic than *ortho*-methoxyphenol.

Question 11.16:

Explain how does the -OH group attached to a carbon of benzene ring activate it towards electrophilic substitution?

Answer 11.16:

The -OH group is an electron-donating group. Thus, it increases the electron density in the benzene ring as shown in the given resonance structure of phenol.



As a result, the benzene ring is activated towards electrophilic substitution.

Question 11.17:

Give equations of the following reactions:

- (i) Oxidation of propan -1-ol with alkaline KMnO₄ solution.
- (ii) Bromine in CS₂ with phenol.
- (iii) Dilute HNO₃ with phenol.
- (iv) Treating phenol with chloroform in presence of aqueous NaOH.

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1. Understanding the eBook Nelson Chemistry 12 Chapter 5 Solutions
 - The Rise of Digital Reading Nelson Chemistry 12 Chapter 5 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 12 Chapter 5 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 12 Chapter 5 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 12 Chapter 5 Solutions
 - Personalized Recommendations
 - Nelson Chemistry 12 Chapter 5 Solutions User Reviews and Ratings
 - Nelson Chemistry 12 Chapter 5 Solutions and Bestseller Lists

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 - Adjustable Fonts and Text Sizes of Nelson Chemistry 12 Chapter 5 Solutions
 - Highlighting and Note-Taking Nelson Chemistry 12 Chapter 5 Solutions
 - Interactive Elements Nelson Chemistry 12 Chapter 5 Solutions
8. Staying Engaged with Nelson Chemistry 12 Chapter 5 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 12 Chapter 5 Solutions
9. Balancing eBooks and Physical Books Nelson Chemistry 12 Chapter 5 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 12 Chapter 5 Solutions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry 12 Chapter 5 Solutions
 - Setting Reading Goals Nelson Chemistry 12 Chapter 5 Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry 12 Chapter 5 Solutions
 - Fact-Checking eBook Content of Nelson Chemistry 12 Chapter 5 Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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