

Pre-lab: Acid-Base Titrations

CHM205 Laboratory

Fall 2019 Section: _____

Grade: _____

Name: _____

A procedure similar to the one in this lab is used to titrate a sodium hydroxide solution against potassium hydrogen phthalate (KHP). The follow data is recorded: For the standard KHP solution, 10.575 g of KHP is dissolved in water, and diluted to 250.00 mL in a volumetric flask; 20.00 mL of the KHP solution are added to an Erlenmeyer flask; a buret is filled with a NaOH solution of unknown concentration to a starting volume of 0.18 mL, and at the end point of the titration the buret reading is 20.83 mL.

1. Calculate the mmoles KHP that were put into the 250.00 mL volumetric flask. Note: KHP's chemical formula is $\text{KHC}_8\text{H}_4\text{O}_4$.
2. Calculate the KHP solution concentration (in molarity units) that was prepared in the 250 mL volumetric flask. Note sig figs on the flask volume.
3. Calculate the mmoles KHP that are in the 20.00 mL aliquot.
4. Write the balanced chemical equation for the reaction of KHP with sodium hydroxide.
5. Calculate the mmoles NaOH that will be needed to reach the titration end point, when the 20.00 mL aliquot is completely titrated.
6. Calculate the volume of NaOH used in the titration.
7. Calculate the concentration (in molarity units) of the NaOH solution.

Acid Base Titrations Pre Lab Answers

Xiang Xie



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Table of Contents Acid Base Titrations Pre Lab Answers

1. Understanding the eBook Acid Base Titrations Pre Lab Answers
 - The Rise of Digital Reading Acid Base Titrations Pre Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titrations Pre Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Titrations Pre Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titrations Pre Lab Answers
 - Personalized Recommendations
 - Acid Base Titrations Pre Lab Answers User Reviews and Ratings
 - Acid Base Titrations Pre Lab Answers and Bestseller Lists
5. Accessing Acid Base Titrations Pre Lab Answers Free and Paid eBooks
 - Acid Base Titrations Pre Lab Answers Public Domain eBooks
 - Acid Base Titrations Pre Lab Answers eBook Subscription Services
 - Acid Base Titrations Pre Lab Answers Budget-Friendly Options

6. Navigating Acid Base Titrations Pre Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid Base Titrations Pre Lab Answers Compatibility with Devices
 - Acid Base Titrations Pre Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid Base Titrations Pre Lab Answers
 - Highlighting and Note-Taking Acid Base Titrations Pre Lab Answers
 - Interactive Elements Acid Base Titrations Pre Lab Answers
8. Staying Engaged with Acid Base Titrations Pre Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Titrations Pre Lab Answers
9. Balancing eBooks and Physical Books Acid Base Titrations Pre Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Titrations Pre Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acid Base Titrations Pre Lab Answers
 - Setting Reading Goals Acid Base Titrations Pre Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid Base Titrations Pre Lab Answers
 - Fact-Checking eBook Content of Acid Base Titrations Pre Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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