

Basic UV-Vis Theory, Concepts and Applications

Introduction

Ultraviolet and visible spectrometers have been in general use for the last 35 years and over this period have become the most important analytical instrument in the modern day laboratory. In many applications other techniques could be employed but none rival UV-Visible spectrometry for its simplicity, versatility, speed, accuracy and cost-effectiveness.

This description outlines the basic principles for those new to UV-Visible spectrometry. It is intended purely as a brief introduction to the technique and it is Thermo Spectronic's policy to continually add to this range of documentation for further details, as they become available.

Definitions and Units

Radiation is a form of energy and we are constantly reminded of its presence via our sense of sight and ability to feel radiant heat. It may be considered in terms of a wave motion where the wavelength, λ , is the distance between two successive peaks. The frequency, ν , is the number of peaks passing a given point per second. These terms are related so that:

$$c = \nu\lambda$$

where c is the velocity of light in a vacuum.



Figure 1 The wavelength λ of electromagnetic radiation

The full electromagnetic radiation spectrum is continuous and each region merges slowly into the next. For spectroscopy purposes, we choose to characterize light in the ultraviolet and visible regions in terms of wavelength expressed in nanometers. Other units which may be encountered, but whose use is now discouraged, are the Angstrom (\AA) and the millimicron ($m\mu$).

$$1\text{nm} = 1m\mu = 10\text{\AA} = 10^{-9}\text{ meters}$$

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