

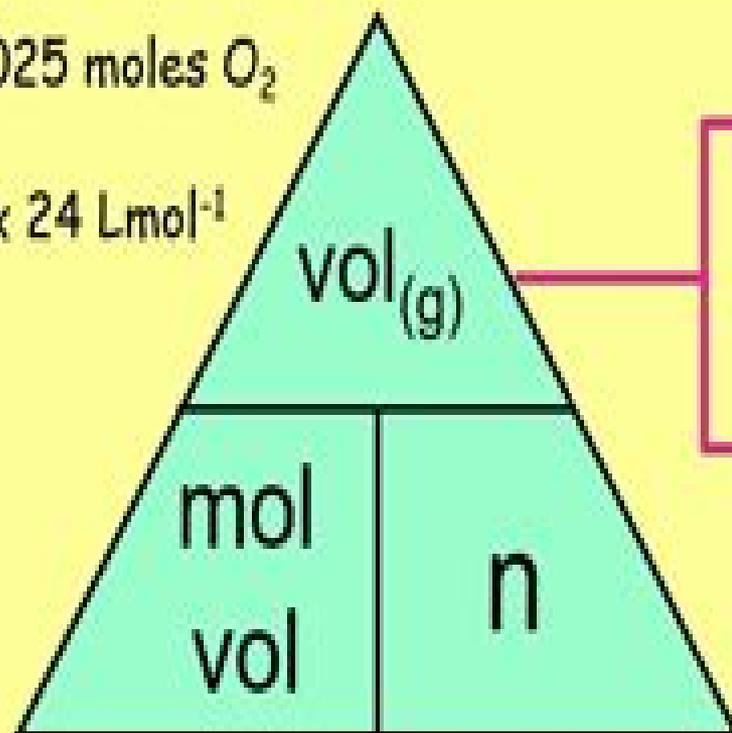
Assume molar volume is 24Lmol^{-1}

Calculate volume of 0.025 moles O_2

$$\text{Vol} = 0.025 \text{ mol} \times 24 \text{ Lmol}^{-1}$$

$$= 0.6 \text{ L}$$

$$= 600 \text{ cm}^3$$



Gas Volume

$$\text{Vol} = n \times \text{mol.vol}$$

Calculations Using Standard Molar Volume

Alexander Silverman



Calculations Using Standard Molar Volume:

Physical Chemistry Robert J. Silbey, Robert A. Alberty, George A. Papadantonakis, Mounji G. Bawendi, 2022-06-15 Ever since *Physical Chemistry* was first published in 1913 it has remained a highly effective and relevant learning tool thanks to the efforts of physical chemists from all over the world Each new edition has benefited from their suggestions and expert advice The result of this remarkable tradition is now in your hands

Chemical Symbolism and Calculations Stanley Wallace Morse, 1953

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Hydrothermal Properties of Materials Vladimir Valyashko, 2008-12-16 Hydrothermal Properties of Materials Experimental Data on Aqueous Phase Equilibria and Solution Properties at Elevated Temperatures and Pressures is designed for any scientists and engineer who

deals with hydrothermal investigations and technologies The book is organized into eight chapters each dealing with a key physical property of behavior of solutions so that a reader can obtain information on hydrothermal experimental methods available experimental data and the main features of properties behavior in a wide range of temperatures and pressures and possible ways of experimental data processing for obtaining the derivative properties

Laboratory Experiments for Foundations of Chemistry Ernest R. Toon,1973

General Chemistry Horace Grove Deming,1925

Beginning Calculations in Physical Chemistry Barry R. Johnson,Stephen K. Scott,1997

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Second Year College Chemistry William Henry Chapin,1922

Salter's Higher Chemistry ,1999 This work provides coverage of the content statements in the arrangements for Higher Chemistry organized by the three units in the course Energy Matters the World of Carbon and Chemical Reactions At the start of each unit students are given guidance on what they need to know and understand

Mechanical Laboratory Methods Julian Chase Smallwood,1924

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