

**Terms to Know:**

- Electrochemistry**—the study of the interchange of chemical and electrical energy
- OIL RIG** – oxidation is loss, reduction is gain (of electrons)
- Oxidation** – the loss of electrons, increase in charge
- Reduction** – the gain of electrons, reduction of charge
- Oxidation number** – the assigned charge on an atom
- Oxidizing agent (OA)** – the species that is reduced and thus *causes* oxidation
- Reducing agent (RA)** – the species that is oxidized and thus *causes* reduction

**ELECTROCHEMISTRY INVOLVES TWO MAIN TYPES OF PROCESSES:**

- A. **Galvanic (voltaic) cells** – which are spontaneous chemical reactions (battery)
- B. **Electrolytic cells** – which are non-spontaneous and require external  $e^-$  source (DC power source)
- C. **BOTH** of these fit into the category entitled **Electrochemical cells**

**GALVANIC CELLS**

Parts of the voltaic or galvanic cell:

- **Anode**--the electrode where oxidation occurs. After a period of time, the anode may appear to become smaller as it falls into solution.
- **Cathode**-- the anode where reduction occurs. After a period of time it may appear larger, due to ions from solution plating onto it.
- **inert electrodes**—used when a gas is involved OR ion to ion involved such as  $Fe^{3+}$  being reduced to  $Fe^{2+}$  rather than  $Fe^0$ . Made of Pt or graphite.
- **Salt bridge** -- a device used to maintain electrical neutrality in a galvanic cell. This may be filled with agar which contains a neutral salt or it may be replaced with a porous cup.
- **Electron flow** -- always from anode to cathode. (through the wire)
- **Standard cell notation (line notation)**.- anode/solution// cathode solution/ cathode Ex.  $Zn/Zn^{2+} (1.0 M) // Cu^{2+} (1.0 M) / Cu$
- **Voltmeter**.- measures the cell potential (*emf*) . Usually is measured in volts.

Balance this re-dox reaction:



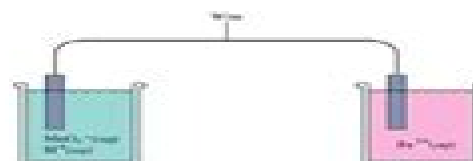
RED:

OA:

OX:

RA

Overall rxn:



# Ap Chemistry Electrochemistry Answers

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